

# Key

## Electron Configuration Evaluation

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium  $1s^2 2s^2 2p^6 3s^1$
- 2) iron  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^10 4p^5$
- 4) barium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^10 4p^6 5s^2$
- 5) neptunium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^10 4p^6 5s^2 4f^14 5d^10 6p^2$   
 $7s^2 5f^9$

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt  $[Ar] 4s^2 3d^7$
- 7) silver  $[Kr] 5s^2 4d^9$
- 8) tellurium  $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium  $[Rn] 7s^2$
- 10) lawrencium  $[Rn] 7s^2 5f^{14}$

Determine what elements are denoted by the following electron configurations:

- 11)  $1s^2 2s^2 2p^6 3s^2 3p^4$  S
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$  Rb
- 13)  $[Kr] 5s^2 4d^{10} 5p^3$  As
- 14)  $[Xe] 6s^2 4f^{14} 5d^0$  Os
- 15)  $[Rn] 7s^2 6d^1 5f^{11}$  Em

Determine which of the following electron configurations are not valid:

- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 5p^3$  Br
- 17)  $1s^2 2s^2 2p^6 3s^2 3d$  Mn
- 18)  $[Ra] 7s^2 5f^2$  Brk
- 19)  $[Kr] 5s^2 4d^{10} 5p^5$  I
- 20)  $[Xe]$  Xe

# Chemistry Guided Practice Problems Answer Key

**WJ Hussar**

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Organic Chemistry, a Guided Inquiry Andrei Straumanis, 2004

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