

11-2 Practice Problems

1. Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:



$$\frac{0.38\text{g Li}_3\text{N}}{1} \times \frac{1\text{mol Li}_3\text{N}}{35\text{g Li}_3\text{N}} \times \frac{3\text{mol LiOH}}{1\text{mol Li}_3\text{N}} \times \frac{24\text{g LiOH}}{1\text{mol LiOH}} = \boxed{0.78\text{g LiOH}}$$

2. What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?



$$\frac{0.29\text{g NaI}}{1} \times \frac{1\text{mol}}{150} \times \frac{2\text{mol NaCl}}{2\text{mol NaI}} \times \frac{58\text{g}}{1\text{mol}} = \boxed{0.112\text{g NaCl}}$$

3. Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:



$$\frac{0.85\text{g C}_4\text{H}_{10}}{1} \times \frac{1\text{mol}}{58\text{g}} \times \frac{8\text{mol CO}_2}{2\text{mol C}_4\text{H}_{10}} \times \frac{44\text{g}}{1\text{mol}} = \boxed{2.58\text{g CO}_2}$$

4. Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:



$$\frac{0.46\text{g Sb}_2\text{O}_3}{1} \times \frac{1\text{mol}}{291.52\text{g}} \times \frac{2\text{mol Sb}}{1\text{mol Sb}_2\text{O}_3} \times \frac{121.76\text{g}}{1\text{mol}} = \boxed{0.38\text{g Sb}}$$

5. What mass of hydrogen peroxide (H_2O_2) must decompose to produce 0.77 g of water?



$$\frac{0.77\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{2\text{mol H}_2\text{O}_2}{2\text{mol H}_2\text{O}} \times \frac{34\text{g}}{1\text{mol}} = \boxed{1.45\text{g H}_2\text{O}_2}$$

6. What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide?



$$\frac{0.69\text{g CO}_2}{1} \times \frac{1\text{mol}}{44\text{g}} \times \frac{2\text{mol CO}}{2\text{mol CO}_2} \times \frac{28\text{g}}{1\text{mol}} = \boxed{0.44\text{g CO}}$$

7. Determine the mass of sodium nitrate produced when 0.73 g of nickel(II) nitrate reacts with sodium hydroxide according to the following equation:



$$\frac{0.73\text{g Ni}(\text{NO}_3)_2}{1} \times \frac{1\text{mol}}{185\text{g}} \times \frac{2\text{mol NaNO}_3}{1\text{mol Ni}(\text{NO}_3)_2} \times \frac{85\text{g}}{1\text{mol}} = \boxed{0.68\text{g NaNO}_3}$$

8. Determine the mass of calcium hydroxide produced when calcium carbide reacts with 0.64 g of water according to the following equation:



$$\frac{0.64\text{g H}_2\text{O}}{1} \times \frac{1\text{mol}}{18\text{g}} \times \frac{1\text{mol Ca}(\text{OH})_2}{2\text{mol H}_2\text{O}} \times \frac{74\text{g}}{1\text{mol}} = \boxed{1.32\text{g Ca}(\text{OH})_2}$$

9. How many grams of ozone (O_3) must decompose to produce 0.87 g of oxygen?



$$\frac{0.87\text{g O}_2}{1} \times \frac{1\text{mol}}{32\text{g O}_2} \times \frac{2\text{mol}}{3\text{mol}} \times \frac{48\text{g O}_3}{1\text{mol}} = \boxed{0.87\text{g O}_3}$$

10. Find the mass of sugar ($\text{C}_6\text{H}_{12}\text{O}_6$) required to produce 1.82 L of carbon dioxide gas at STP from the reaction described by the following equation:



11. How many liters of oxygen are necessary for the combustion of 425 g of sulfur, assuming that the reaction occurs at STP? The balanced equation is $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$.

12. Find the mass of benzene (C_6H_6) required to produce 2.66 L of carbon dioxide gas at STP from the reaction described by the following equation:



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Susan McMurry, John McMurry



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Adkins,2003

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McMurry,Susan McMurry,1986

Principles and Practices of Automatic Process Control Carlos A. Smith,Armando B.

Corripio,2005-08-05 Highly practical and applied this Third Edition of Smith and Corripio's Principles and Practice of Automatic Process Control continues to present all the necessary theory for the successful practice of automatic process control The authors discuss both introductory and advanced control strategies and show how to apply those strategies in industrial examples drawn from their own professional practice The strengths of the book are its simplicity excellent examples practical approach real case studies and focus on Chemical Engineering processes More than any other textbook in the field Smith Corripio prepares a student for use of process control in a manufacturing setting Course Hierarchy Course is called Process Control Senior level course Same course as Seborg but Smith is considered more accessible Electroactive Polymer Electrochemistry Michael E.G. Lyons,2013-03-09 A comprehensive review of the current state of the theoretical development in this important area of potential application of conducting polymers and is very timely The editor author is to be congratulated for his marathon efforts and the production of a significant contribution to the literature TRIP This three part series provides undergraduate and graduate students in electrochemistry and materials science with a broad understanding of electroactive polymers In Part I renowned scientists examine the fundamental principles underlying electrochemical behavior of electroactive polymer materials Contributors focus on the fundamentals of charge percolation and conductivity behavior associated with the membrane properties of electroactive polymer films Part I also includes coverage of the phenomenon of heterogeneous redox catalysis at electroactive polymer modified electrodes

Introductory Chemistry Lillian Hoagland Meyer,1964

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