

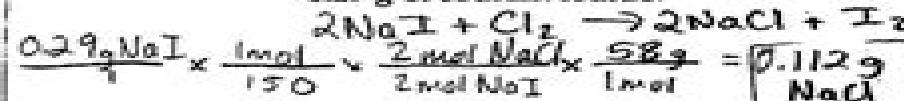
11-2 Practice Problems

1. Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:

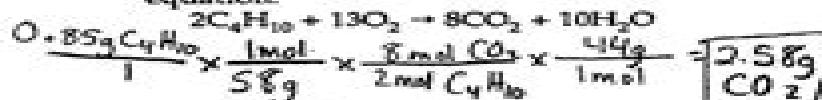


$$\frac{0.38 \text{ g Li}_3\text{N}}{1} \times \frac{1 \text{ mol Li}_3\text{N}}{35.5 \text{ g Li}_3\text{N}} \times \frac{3 \text{ mol LiOH}}{1 \text{ mol Li}_3\text{N}} \times \frac{24.0 \text{ g LiOH}}{1 \text{ mol LiOH}} = 0.78 \text{ g LiOH}$$

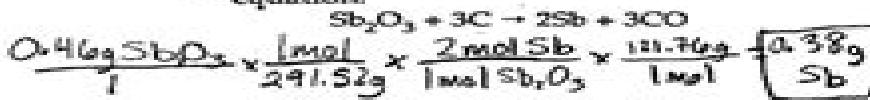
2. What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?



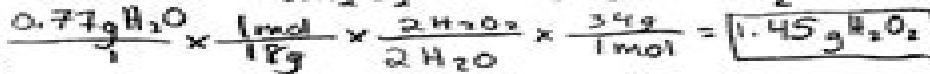
3. Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:



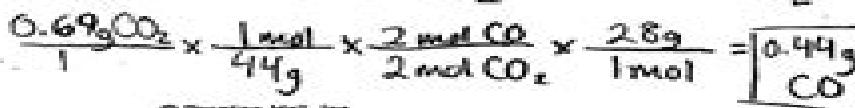
4. Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:



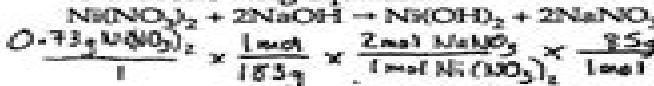
5. What mass of hydrogen peroxide (H_2O_2) must decompose to produce 0.77 g of water?



6. What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide?

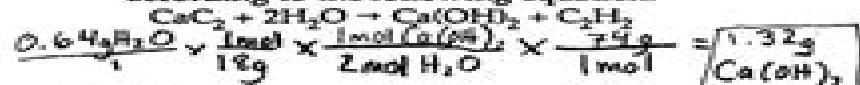


7. Determine the mass of sodium nitrate produced when 0.73 g of nickel(II) nitrate reacts with sodium hydroxide according to the following equation:



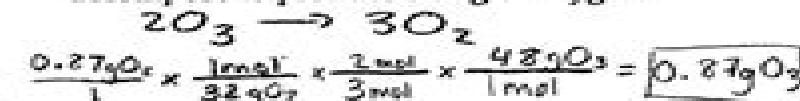
0.68 g
NaNO₃

8. Determine the mass of calcium hydroxide produced when calcium carbide reacts with 0.64 g of water according to the following equation:



1.32 g
Ca(OH)₂

9. How many grams of ozone (O_3) must decompose to produce 0.87 g of oxygen?



10. Find the mass of sugar ($\text{C}_6\text{H}_{12}\text{O}_6$) required to produce 1.82 L of carbon dioxide gas at STP from the reaction described by the following equation:



11. How many liters of oxygen are necessary for the combustion of 425 g of sulfur, assuming that the reaction occurs at STP? The balanced equation is $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$.

12. Find the mass of benzene (C_6H_6) required to produce 2.66 L of carbon dioxide gas at STP from the reaction described by the following equation:



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Susan McMurry, John McMurry

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Prentice Hall Grammar Jeanette Adkins,2003

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