

## **Chapter 10 Review: Chemical Quantities & Calculations**

*(aka – mole & volume calculations, % composition, empirical & molecular formulas)*

For the test, be sure that you are familiar, and able to complete, the problems below as well as the problems found on the following worksheets: "Molecular Weight and Mole Calculations," "Percent Comp. & Chemical Formula Calculations" and "Chapter 10 Sample Problems." Check the class website for answers to the review worksheet.

1. The air around you is made up of 20.95% oxygen. How many oxygen molecules are in 75.0 grams of air? The density of air is 1.19 g/L. *assume STP.*
2. How much helium (*grams*) would be needed to inflate a balloon to a volume of 5.50L? *assume STP.*
3. The molecular formula of an antibacterial drug is  $C_{17}H_{18}FN_2O_5$ . How many fluorine atoms are in 150mg of this drug?
4. A molecule is made up of 40.0% carbon, 6.67% hydrogen, and an unknown amount of oxygen. What is the molecular formula of the molecule if its molar mass is 90.0g/mole?
5. What is the empirical formula for a compound made up of 71.72% chlorine, 16.16% oxygen and 12.12% carbon? Calculate the molecular formula for the compound if it has a molar mass of 296.7g/mol.
6. How many molecules are in 8.42 Kilograms of  $C_2H_4$ ?

# Chapter 10 Chemical Quantities Guided Practice Problems Answers

**Michael Rajnik**



## **Chapter 10 Chemical Quantities Guided Practice Problems Answers:**

*Study Guide/Selected Solutions Manual* Julie R. Frentrup, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

*Chemistry* John S. Phillips, Cheryl Wistrom, 2000      **Study guide to accompany chemistry** Karen Timberlakes, 1983

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students      **Study Guide with Selected Solutions** Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems      *Study Guide to Accompany Principles of Microeconomics* David R. Hakes, 2000-08 To accomplish your course goals use this study guide to enhance your understanding of the text content and to be better prepared for quizzes and tests This convenient manual helps you assimilate and master the information encountered in the text through the use of practice exercises and applications comprehensive review tools and additional helpful resources      **Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed** Peter P. Berlow, 1978      Study Guide for Chemical Principles Thomas Elliott Taylor, 1979      **Study Guide [to Accompany] General Chemistry** Dixie Goss, 2002      **The Pharmaceutical Journal and Pharmacist**, 1909      **Pharmaceutical Journal**, 1909      **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science**, 2003-11 Prentice Hall Physical Science Concepts in Action

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