

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

- An apple pie needs 10 large apples, 2 cups of flour, and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed for one pie?
- Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, 2KClO_3 . Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
- Using the equation from problem 2, how many moles of oxygen are produced from 1.0 moles of potassium chlorate?
- Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 2.0×10^{24} molecules of oxygen? How many moles of water are produced from 22.0×10^{24} molecules of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

- Calculate the number of moles of hydrogen chloride produced from 30 moles of hydrochloric acid.



- Calculate the number of moles of chlorine needed to form 1.0 moles of iron(III) chloride.



- Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



- Calculate the mass, in grams, of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



- Calculate the mass, in grams, needed to react with chlorine to produce 164 g of silver chloride. (Refer to Table 1 for the molar mass of each element.)

- How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



- Calculate the number of moles of oxygen gas needed to produce 4.00 liters of nitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



Chapter 12 Stoichiometry Reading Guide

**Daniel L. Reger, Edward E.
Mercer, Robert S. Boikess**

Chapter 12 Stoichiometry Reading Guide:

CliffsAP Chemistry, 4th Edition Bobrow Test Preparation Services, 2011-09-26 Your complete guide to a higher score on the AP Chemistry exam Why CliffsAP Guides Go with the name you know and trust Get the information you need fast Written by test prep specialists Contents include Introduction overview of the test and how it is scored proven strategies for each type of question Review of topics tested atom periodic table bonding geometry hybridization stoichiometry gases liquids and solids thermodynamics solutions equilibrium acids and bases kinetics redox nuclear chemistry organic chemistry and writing reactions The Labs feature 20 multiple choice questions multiple free response questions on each topic with answers on each topic with answers and explanations scoring rubrics and 2 full length practice exams Structured like the actual exam Complete with answers and explanations AP is a registered trademark of the College Board which was not involved in the production of and does not endorse this product

Teaching Science for Understanding James Joseph Gallagher, 2007

Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities

CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book

Study Guide for Chemistry by Steven S. Zumdahl Martha B. Barrett, 1986

Student Study Guide to accompany Chemistry Martin Silberberg, Libby Weberg, 2005-01-06

Essential Concepts of Chemistry Study Guide James R. Braun, Sherman, 1999

Study Guide

[to Accompany] General Chemistry Dixie Goss, 2002

Student Study Guide to Accompany Petrucci's General Chemistry

3rd. Ed Robert K. Wismer, 1982

Student Study Guide Fundamentals of Chemistry by David E. Goldberg Margaret G. Kimble, 1994

Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995

The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2005-02-02

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Study Guide for Chemistry, Third Edition [by] **Steven S. Zumdahl** Paul B. Kelter,1993 *Student Study Guide to Accompany Petrucci's General Chemistry* Robert K. Wissner,1985 **Fundamentals of Chemistry, Study Guide** James E. Brady,John R. Holum,1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

Student Study Guide for Introduction to Chemistry John Dickson,1979-01-25 **Study Guide General Chemistry, Second Edition, Becker/Wentworth** Becker,Benjamin L. Carroll,1980 **Study Guide/Selected Solutions Manual** Julie R. Frentrup,2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book [Introduction to Chemistry, Study Guide](#) T. R. Dickson,1999-07-07 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and chemical vision and refining problem solving skills Many chapters in this edition have been rewritten and rearranged to vitalize the topics and to include interesting examples analogies and images **Study Guide and Glossary to Accompany Chemical Principles** Daniel L. Reger,Edward E. Mercer,Robert S. Boikess,1978 *Study guide to accompany Drew H. Wolfe: General, organic and biological chemistry* Daniel C. Pantaleo,1986

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