

CHAPTER 15 **STUDY GUIDE****Energy and Chemical Change****Section 15.1 Energy***In your textbook, read about the nature of energy.***In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.**

1. Energy is the ability to do work or produce heat.
2. The law of conservation of energy states that energy *can be* created and destroyed.
3. Chemical potential energy is energy stored in a substance because of its *composition*.
4. *Heat* is a form of energy that flows from a warmer object to a cooler object.
5. A caloric is the amount of energy required to raise the temperature of *one gram* of pure water by one degree Celsius.
6. A *calorie* is the SI unit of heat and energy.
7. The specific heat of a substance is the amount of heat required to raise the temperature of *one gram* of that substance by one degree Celsius.
8. *Kinetic energy* is energy of motion.
9. Chemicals participating in a chemical reaction contain only *potential energy*.
10. One nutritional Calorie is equal to *100 calories*.
11. One caloric equals *4.184 joules*.
12. When a fuel is burned, some of its *chemical potential energy* is lost as heat.
13. To convert kilojoules to joules, *divide* the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is $4.184 \text{ J(g}^{\circ}\text{C)}$.

Chapter 15 Chemistry Study Guide Answers

Margaret G. Kimble

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Wastewater Operator Certification Study Guide John Giorgi,2011-01-12 Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e Kenneth W. Raymond,2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **Study Guide to Accompany Basics for Chemistry** Martha Mackin,2012-12-02

Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students Student Study Guide Fundamentals of Chemistry by David E. Goldberg Margaret G. Kimble,1994

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Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding. Every morning many of us are energized by a cup of coffee. Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans. Now that's a real chemical understanding. Olmsted and Williams' Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist. This text helps you understand how chemical phenomena are governed by what happens at the molecular level, apply critical thinking skills to chemical concepts and problems, and master the basic mathematical techniques needed for quantitative reasoning. You'll see the world as chemists do and learn to appreciate the chemical processes all around us. A Fourth Edition with a lot of new perks. Revisions include a new early energy chapter, revised coverage of bonding, expanded coverage of intermolecular forces, and increased coverage of multiple equilibria including polyprotic acids. New pedagogy strengthens students' critical thinking and problem solving skills. Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills, concepts, equations, and terms. eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text, links between problems and relevant sections in the online text, practice quizzes, the Visual Tutor, Interactive LearningWare problems and lab demos, as well as homework management and presentation features for instructors.

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