

Study Guide: Chapter 15 Solutions (pg 1 of 2)

It is important to know how much solid is dissolved in a liquid. Telling someone that you added a little sugar to their ice tea does not give them much of a clue about how sweet the tea will be. A little sugar to you might mean a lot of sugar to someone else. In this chapter you learned ways to express concentration so that you knew just how much dissolved solid is present in a given volume of liquid.

Many common ionic and nonionic substances dissolve in water. When an ionic substance dissolves, it breaks apart into ions. For example, when potassium sulfate, K_2SO_4 , dissolves, it forms K^+ and SO_4^{2-} ions. To understand why a crystal of potassium sulfate breaks apart into ions, we need to consider the nature of the water molecule. Water molecules are polar; that is, one end of the molecule has a partial positive charge, and the other end has a partial negative charge. The positive potassium ion is attracted to the partial negative charge on the water molecule, and the negative sulfate ion is attracted to the partial negative charge on the water molecule. The K_2SO_4 crystal is pulled apart by the polarity of the water molecule. Ions in solution are surrounded by oppositely charged ends of water molecules.

Nonionic compounds such as ethyl alcohol, C_2H_5OH , also dissolve in water. The O-H of ethyl alcohol is polar, just as the O-H on the water molecule is. This means that alcohol molecules have a negative end and a positive end just as water molecules do. Alcohol molecules are attracted by water molecules and are dissolved in them.

Molecules that are not water soluble do not have positive or negative ends to be attracted to water.

One way of describing the composition of a solution is mass percent. The mass percent of a solution is the mass of solute dissolved by the total mass of the solute plus solvent, multiplied by 100%.

$$\text{mass percent} = (\text{mass of solute}/\text{mass of solution}) \times 100\%$$

However, the mass percent of a solution is inconvenient to use when the solvent is liquid. It is much more convenient to measure the volume of a liquid than it is to measure its mass. The most often used indication of amount of solute in a given volume of solution is molarity. Molarity is equal to the number of moles of solute per volume of solution. It is often abbreviated as $M = \text{mol/L}$.

$$\text{Molarity} = (\text{moles of solute}/\text{liters of solution})$$

When the molarity of a solution is calculated, it is assumed that the solute is in the form it would have been in before it dissolved. One mole of Na_2SO_4 , when added to water, produces two moles of Na^+ ions and one mole of SO_4^{2-} ions.



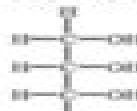
Compounds that do not form ions have the same molar concentration before and after they dissolve in water. Sucrose, table sugar, dissolves in water but does not form ions. 1 mol of sucrose in 1 liter of solution produces a 1 M solution of sucrose.

Dilution is the process of adding more solvent to a solution. When we prepare a dilute solution, we are measuring a quantity of a stock solution (relatively highly concentrated solution) and adding it to water. The amount of solute in the measured portion of stock solution is the same as the amount of solute in the more dilute solution. The only thing which has changed is the volume of the solution. We have decreased the concentration of the solution by increasing the volume, but the amount of solutes stays the same.

In Chapter 9, you learned to solve stoichiometry problems. From the balanced equation you could answer questions about the quantity of reactant required or the quantity of product produced. The same principles are used here, but the reactions occur in solutions. In Chapter 9, we used molar mass to convert mass to moles. Here we use molarity to convert volume of solution to moles. Use the steps in Section 15.6 of your text.

Learning Review:

1. 150 mL of ethyl alcohol is mixed with 1 L of water. Which is the solute, ethyl alcohol or water?
2. Which of the molecules below would you predict to be soluble in water?



- a. $CH_3-CH_2-CH_3$
- b. K_2SO_4
- c. Na_2SO_4

Chapter 15 Solutions Study Guide

Don Miller

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Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder s tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

Study Guide/Selected Solutions Manual Julie R. Frentrup, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

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universities throughout the United States and Canada and is an international bestseller from the United Kingdom to the Pacific Rim In this edition McMurry continues to do what he does best focus on the important material of the course and explain it in a concise clear way Student Study Guide with Selected Solutions Barbara Snyder,1996-08 Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e Kenneth W. Raymond,2005-10-07

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