

CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Covalent Bonding

Section 11 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	covalent	sigma bond	double bond	p bond
When sharing of electrons occurs, the electrons between atoms that results is called	(a) covalent bond	When such as the bond is formed by electrons between two atoms		

(b) **σ bond** _____ When such as the bond is formed by electrons between two atoms, and the position is (b) **σ bond** _____ When two or more atoms bond by means of electrons sharing the resulting patch is called an (c) **σ molecule** _____ If the electrons shared are located between the two atoms, the electrons are called an (d) **σ bond** _____ If the electrons shared are located between the two atoms, the position of which, the electrons is called an (e) **σ bond** _____

In your textbook, read about single and double bonds, and bond strength.

Circle the letter of the theory that best completes the statement in answer to the question.

6. In what form do electrons in two hydrogen atoms, and oxygen atom(s) move?

a. as single atoms
 b. as molecules containing two atoms
 c. as molecules containing four atoms
 d. as molecules containing two bonds

7. How many electrons are shared in a double covalent bond?

a. one
 b. two
 c. four
 d. six

8. Bond length is the distance between

a. two molecules of the same substance
 b. the electrons in two different atoms
 c. the nuclei of two different atoms
 d. the nuclei of two identical atoms

9. What is the following relationship referring to bond length in generally untrue?

a. As the size of the bond, the length of the bond
 b. the shorter the bond, the weaker the bond
 c. the shorter the bond, the lower the electrons in it
 d. the shorter the bond, the larger the bond length

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Section 12 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write true or false.

1. Many molecular compounds are generally composed of a metal and a nonmetal.

2. The general formula for the formula of a binary compound is named using the suffix -ide.

3. The prefix from indicates three atoms.

4. The prefix from indicates four atoms.

5. In naming the first element in a formula, the prefix comes in last.

6. For binary acids, the hydrogen part of the compound is named using the prefix hydro-.

7. An element contains only two elements.

8. If the name of the element in a binary acid is on the end of the name, remove the suffix -ide.

In your textbook, read about naming molecular compounds and acids.

For each item in Column A, write the letter of the matching item in Column B.

Column A

1. **H₂O**
 2. **Na₂O**
 3. **Na₂CO₃**
 4. **Ca₃(PO₄)₂**
 5. **Na₂SiO₃**
 6. **Ca(OH)₂**
 7. **Na₂SO₄**
 8. **Na₂CO₃**

Column B

a. hydronium acid
 b. aluminum hydroxide
 c. carbonic acid
 d. calcium carbonate
 e. calcium hydroxide
 f. calcium oxide
 g. calcium phosphate
 h. calcium silicate
 i. sodium acid
 j. sodium carbonate
 k. sodium hydroxide
 l. sodium oxide

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**Clifford Thomas Morgan, Richard
Austin King, James W. Rosen**

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