

CHAPTER 9 REVIEW**Stoichiometry****SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ The coefficients in a chemical equation represent the
 - (a) masses in grams of all reactants and products.
 - (b) relative number of moles of reactants and products.
 - (c) number of atoms of each element in each compound in a reaction.
 - (d) number of valence electrons involved in a reaction.
2. _____ Which of the following would not be studied within the topic of stoichiometry?
 - (a) the mole ratio of Al to Cl in the compound aluminum chloride
 - (b) the mass of carbon produced when a known mass of sucrose decomposes
 - (c) the number of moles of hydrogen that will react with a known quantity of oxygen
 - (d) the amount of energy required to break the ionic bonds in CaF_2
3. _____ A balanced chemical equation allows you to determine the
 - (a) mole ratio of any two substances in the reaction.
 - (b) energy released in the reaction.
 - (c) electron configuration of all elements in the reaction.
 - (d) reaction mechanism involved in the reaction.
4. _____ The relative number of moles of hydrogen to moles of oxygen that react to form water represents a(n)
 - (a) reaction sequence.
 - (b) bond energy.
 - (c) mole ratio.
 - (d) element proportion.

5. Given the reaction represented by the following unbalanced equation:

$$\text{N}_2\text{O}(g) + \text{O}_2(g) \rightarrow \text{NO}_2(g)$$

- a. Balance the equation.

- b. What is the mole ratio of NO_2 to O_2 ?
- c. If 20.0 mol of NO_2 form, how many moles of O_2 must have been consumed?
- d. Twice as many moles of NO_2 form as moles of N_2O are consumed. True or False?
- e. Twice as many grams of NO_2 form as grams of N_2O are consumed. True or False?

Chapter 9 Reading Guide Stoichiometry

Leo J. Malone

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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills,Amina El-Ashmawy,2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity **Modern Ceramic Engineering** David W. Richerson,2005-11-04 Ceramic materials have proven increasingly important in industry and in the fields of electronics communications optics transportation medicine energy conversion and pollution control aerospace construction and recreation Professionals in these fields often require an improved understanding of the specific ceramics materials they are using **Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed** Robert K. Wismer,1982 **Study Guide and Glossary to Accompany Chemical Principles** Daniel L. Reger,Edward E. Mercer,Robert S. Boikess,1978 **Basic Concepts of Chemistry, Study Guide** Leo J. Malone,1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples

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