

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent
insoluble	solid	solvent

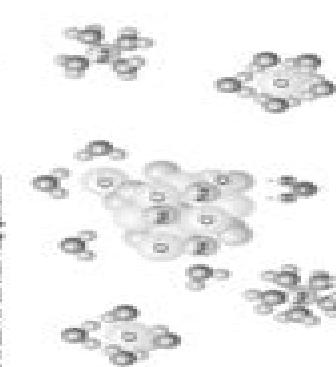
After reading (1) _____ solution _____ of solute is dissolved in solvent.

(2) _____ solute _____ is dissolved in the (3) _____ solvent _____.

(4) _____ solute _____ is dissolved in a solvent, or (5) _____ is in (6) _____ solution. A substance that does not dissolve is (7) _____ insoluble _____ in (8) _____ liquid _____ solution or the most common type of solution. If one liquid is soluble in another liquid, such as water and oil, the two liquids are (9) _____ miscible _____ because if one liquid is insoluble in another, the liquids are (10) _____ immiscible _____.

Read about solubility in aqueous solution in your textbook.

Use the diagram to answer the following questions.



1. Indicate in solute in which the solute is in a (a) (b) (c) (d)

(a) solution is the process of surrounding solute particles with solvent particles to form a solution.

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Chemistry: Matter and Change • Chapter 15

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Section 15.2 Solubility**STUDY GUIDE FOR CONTENT MASTERY**

16. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a glass cup, add the rest of solute may be dissolved.

Starting at the bottom of the solution, breaking the solution into smaller pieces, and heating the solvent.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvents, with tips or jobs.

True _____ 16. The overall energy change that occurs when a solute dissolves is heat of solution.

False _____ 17. Solubility is a property of the solute, amount of water the solute is in a given amount of solvent at a specified temperature and pressure.

False _____ 18. Solubility continues to heat as the solution heat is lost due to the evaporation of water.

True _____ 19. In a saturated solution, solute and crystallization are in equilibrium.

True _____ 20. Additional solute can be dissolved in an unsaturated solution.

True _____ 21. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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DP Hallahan

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