

# Equilibrium: Assessment Review

Answers

Skills - The student will be able to draw and explain how a catalyst changes a reaction rate graph and discuss what changes that disrupt an equilibrium system.

2. Using the reaction diagram to the right, representing the reaction  $A + B \rightleftharpoons C + D$ , draw a new curve that shows how a catalyst would affect the reaction pathway.

3. Place a checkmark next to the actions that would most likely **accelerate** the reaction rate:

- Increasing the temp of the system
- Increasing the concentration of reactants present
- Adding a solid in to the reaction
- Increasing a enzyme catalyst



4. Explain why all reactions have activation energy, using your knowledge of collision theory.

5. Explain why different reactions have different activation energies.

6. Describe how the activation energy of a reaction affects the overall rate of the chemical reaction differently than a change in concentration, pressure, or temp.

7. A rule of thumb used by most chemists is that the rate of a chemical reaction can be doubled by increasing the reaction temperature by ten degrees Celsius. Explain this direct increase in reaction rate using your knowledge of the collision model.

Skills - The student will be able to write the equilibrium expression, given a reaction, draw a reaction coordinate diagram, and calculate the amounts of reactants and products at equilibrium. A reaction may require the step of calculating equilibrium fractions, the amounts of products and reactants that will change. This process is called stoichiometry or equilibrium.

A reaction that has reached equilibrium can be described with a ratio between the reactants and products. This is called the equilibrium constant, calculated by the concentrations of reactants, including the equilibrium constant,  $K_c$ , where existing equilibrium concentrations, only equilibrium and general equilibrium are described.

Balance the following equations & write the equilibrium expression, including the value for  $K_c$ .



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