

# Equilibrium: Assessment Review

Name \_\_\_\_\_

**SKILL -** The student will be able to draw and explain how a catalyst changes a reaction rate graph and discuss other changes that disrupt an equilibrium system.

1. Using the reaction diagram to the right, representing the reaction:  $A + B \rightarrow C + D$  draw a new curve that shows how a catalyst would affect the reaction pathway.

2. Place a checkmark next to the actions that would most likely decrease the reaction rate.

- ☐ Lowering the temp of the system
- ☐ Increasing the concentration of substances present
- ☐ Grinding a solid in to fine particles
- ☐ Removing a enzyme catalyst

3. Explain why all reactions have activation energy, using your knowledge of collision model.

4. Explain why different reactions have different activation energies.

5. Describe how the activation energy of a reaction affects the overall rate of the chemical reaction differently than a change in concentration, pressure, or temp.

6. A rule of thumb used by most chemists is that the rate of a chemical reaction can be doubled by increasing the reaction temperature by ten degrees Celsius. Explain this drastic increase in reaction rate using your knowledge of the collision model.



**SKILL -** The student will be able to write the equilibrium expression, given a reaction.

When a reaction reaches equilibrium the reverse reaction forms reactants at the same rate that the forward reaction produces products. A reaction may appear to stop when it reaches equilibrium because the amount of products and reactants does not change. This process is called dynamic equilibrium.

A reaction that has reached equilibrium can be described with a ratio known as the equilibrium expression. This expression relates the concentration of products divided by the concentration of reactants, raising for the equilibrium constant,  $K_c$ . When writing equilibrium expressions, only aqueous and gaseous substances are included.

Balance the following equations & write the equilibrium expressions needed to solve for  $K_c$ .



# Chemical Equilibrium Review Answer Key

**The Princeton Review**



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**Table of Contents Chemical Equilibrium Review Answer Key**

1. Understanding the eBook Chemical Equilibrium Review Answer Key
  - The Rise of Digital Reading Chemical Equilibrium Review Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equilibrium Review Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Equilibrium Review Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Equilibrium Review Answer Key
  - Personalized Recommendations
  - Chemical Equilibrium Review Answer Key User Reviews and Ratings
  - Chemical Equilibrium Review Answer Key and Bestseller Lists
5. Accessing Chemical Equilibrium Review Answer Key Free and Paid eBooks
  - Chemical Equilibrium Review Answer Key Public Domain eBooks
  - Chemical Equilibrium Review Answer Key eBook Subscription Services
  - Chemical Equilibrium Review Answer Key Budget-Friendly Options
6. Navigating Chemical Equilibrium Review Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemical Equilibrium Review Answer Key Compatibility with Devices
  - Chemical Equilibrium Review Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Equilibrium Review Answer Key
  - Highlighting and Note-Taking Chemical Equilibrium Review Answer Key
  - Interactive Elements Chemical Equilibrium Review Answer Key
8. Staying Engaged with Chemical Equilibrium Review Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemical Equilibrium Review Answer Key
- 9. Balancing eBooks and Physical Books Chemical Equilibrium Review Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Equilibrium Review Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Equilibrium Review Answer Key
  - Setting Reading Goals Chemical Equilibrium Review Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Equilibrium Review Answer Key
  - Fact-Checking eBook Content of Chemical Equilibrium Review Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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