

Lab Report Experiment 11: Chemical Kinetics

Data Table I (Rate vs Concentration) The temperature is 45 °C.

Trial	H ₂ O (mL)	0.010 M H ₂ O ₂ (mL)	1.0 M HCl (mL)	0.050 M KI (mL)	Time (sec)	Rate (1/sec)
1	60.0	10.0	20.0	10.0	45	
2	50.0	20.0	20.0	10.0	23	
3	40.0	10.0	40.0	10.0	45	
4	50.0	10.0	20.0	20.0	24	

Data Table II (Rate vs Temperature)

Trial	H ₂ O (mL)	0.010 M H ₂ O ₂ (mL)	1.0 M HCl (mL)	0.050 M KI (mL)	Temp °C	Time (sec)	Rate (1/sec)
1	60.0	10.0	20.0	10.0	5	67	
2	60.0	10.0	20.0	10.0	15	60	
3	60.0	10.0	20.0	10.0	25	55	
4	60.0	10.0	20.0	10.0	35	52	
5	60.0	10.0	20.0	10.0	45	45	

Analysis:

1. What happened to the rate when the amount of H₂O₂ was doubled? What is the order of the reaction with respect to H₂O₂?
2. What happened to the rate when the amount of HCl was doubled? What is the order of the reaction with respect to HCl?

Chemical Kinetics Lab Report

David Garvin

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Reaction Rates for High-temperature Air with Carbon and Sodium Impurities Mina L. Carnicom, 1968 The values used by a number of investigators for the rate constants of high temperature greater than or equal to 1000 K homogeneous gaseous reactions involving species of the elements nitrogen oxygen carbon and sodium have been compiled and are presented in tabular form Included are reactions involving neutral species charged species free electrons some species in excited electronic or vibrational states and radiative processes SCR ,1968 Scientific and Technical Aerospace Reports ,1963

Integrated Approach to Coordination Chemistry Rosemary A. Marusak, Kate Doan, Scott D. Cummings, 2007-03-07
Coordination chemistry is the study of compounds formed between metal ions and other neutral or negatively charged molecules This book offers a series of investigative inorganic laboratories approached through systematic coordination chemistry It not only highlights the key fundamental components of the coordination chemistry field it also exemplifies the historical development of concepts in the field In order to graduate as a chemistry major that fills the requirements of the American Chemical Society a student needs to take a laboratory course in inorganic chemistry Most professors who teach and inorganic chemistry laboratory prefer to emphasize coordination chemistry rather than attempting to cover all aspects of inorganic chemistry because it keeps the students focused on a cohesive part of inorganic chemistry which has applications in medicine the environment molecular biology organic synthesis and inorganic materials **Ninth Symposium**

(International) on Combustion Sam Stuart, 2013-09-11 Ninth Symposium International on Combustion covers the proceedings of the Ninth Symposium International on Combustion held at Cornell University in Ithaca New York on August 27 to September 1 1962 under the auspices of the Combustion Institute The book focuses on the processes and reactions involved in combustion The selection first offers information on flame strength of propane oxygen flames at low pressures in turbulent flow and mixing and flow in ducted turbulent jets Topics include radial profile of the jetting velocity radial growth of the jet and mixing zones of a ducted jet The text then elaborates on turbulent flame studies in two dimensional open burners turbulent mass transfer and rates of combustion in confined turbulent flames and flame stabilization in a boundary layer The publication examines the theoretical study of properties of laminar steady state flames as a function of properties of their chemical components and spectra of alkali metal organic halide flames The text then takes a look at the thermal radiation theory for plane flame propagation in coal dust clouds flame characteristics of the diborane hydrazine system and studies of the combustion of dimethyl hydrazine and related compounds The selection is a dependable reference for readers interested in the processes and reactions involved in combustion Combustion & Emission Formation Process in Diesel Engines ,2003 **Journal de physique** ,2002 *A Compendium of Evaluated and Estimated Rate Coefficients* David Garvin, 1968 *Energy Research Abstracts* ,1977 Semiannual with semiannual and annual indexes References to all scientific and technical literature coming from DOE its laboratories energy centers and contractors Includes all works

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