

Chemistry 110 Study Guide Chapters 55-57

Chapter 55

Key Concepts:

1. The electronic structure of atoms determines their chemical properties.
2. The types of elements are determined by the number of protons in the nucleus.
3. The number of protons in an element's nucleus determines how strong the atom's nucleus is held together (termed "atomic number").
4. The types of elements are determined by the number of electrons in the outermost energy level of the atom (termed "atomic number").
5. Electrons in atoms are arranged in shells around the nucleus.
6. The shells are filled in order of increasing energy.
7. The first shell can hold a maximum of 2 electrons, the second shell can hold a maximum of 8 electrons, and so on.
8. The outermost shell is called the valence shell.
9. The valence shell determines how an atom will interact with other atoms.
10. Atoms interact with each other in order to achieve a full valence shell.
11. The types of elements are determined by the number of valence electrons in the outermost shell of the atom.
12. The number of valence electrons determines how an atom will interact with other atoms.
13. Atoms with the same number of valence electrons will interact with each other in the same way.

Important Events:

1. The discovery of the electron in 1897 by J. J. Thomson.
2. The discovery of the nucleus in 1911 by Ernest Rutherford.
3. The discovery of the proton in 1918 by Ernest Rutherford.
4. The discovery of the neutron in 1932 by James Chadwick.
5. The development of the periodic table of the elements by Dmitri Mendeleev in 1869.

Notable Figures:

1. J. J. Thomson - discovered the electron in 1897.
2. Ernest Rutherford - discovered the nucleus in 1911 and the proton in 1918.

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John Blair Russell

Chemistry 110 Study Guide:

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon: buckyballs, graphite, and diamond are illustrated at the left as is the molecule methane CH₄ from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

Study Guide & Solutions Manual to Accompany Organic Chemistry, Third Edition G. Marc Loudon, Joseph G. Stowell, 1995 *Laboratory Directions and Study Questions in Inorganic Chemistry* Alexander Silverman, 1919 *Organic Chemistry, 4e Student Solution Manual and Study Guide* David R. Klein, 2021-01-07 Success in organic chemistry requires mastery in two core aspects: fundamental concepts and the skills needed to apply those concepts and solve problems. With *Organic Chemistry Student Solution Manual and Study Guide* 4th Edition, students can learn to become proficient at approaching new situations methodically based on a repertoire of skills. These skills are vital for successful problem solving in organic chemistry.

Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed Robert K. Wismer, 1982 *Student Study Guide to Accompany Petrucci's General Chemistry* Robert K. Wismer, 1985

Organic Chemistry, 5e Student Study Guide and Solutions Manual David R. Klein, Laurie S. Starkey, 2025-03-18 Success in organic chemistry requires mastery in two core aspects: fundamental concepts and the skills needed to apply those concepts and solve problems. With *Organic Chemistry Student Study Guide and Solutions Manual* 5th Edition, students can learn to become proficient at approaching new situations methodically based on a repertoire of skills. These skills are vital for successful problem solving in organic chemistry.

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Introduction to Chemistry, Study Guide T. R. Dickson, 1999-07-07 Teaches chemistry by offering a dynamic, provocative and relevant view of the topic and its importance to society and our daily lives. Three themes are stressed throughout the text: developing chemical thinking and chemical vision and refining problem solving skills. Many chapters in this edition have been rewritten and rearranged to vitalize the topics and to include interesting examples, analogies and images.

Journal of Chemical Education, 1928 *Manual of Chemistry* William Simon, Daniel Base, 1912 **Study Guide to Accompany Sienko and Plane['s] Chemistry, Principles and Properties** John Blair Russell, 1967 **Study Guide**

and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition Susan McMurry,2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines **Study Guide [to] Chemistry, a Systematic Approach [by] Sisler, Dresden, and Mooney** William H. Myers,William T. Mooney,1980 **Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl** Paul B. Kelter,1993

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