

Types of Chemical Reactions Worksheet #1

H/Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1) $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ **Synthesis**
- 2) $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$ **Single Replacement**
- 3) $3\text{KOH} + \text{AlCl}_3 \rightarrow \text{Al}(\text{OH})_3 + 3\text{KCl}$ **Double Replacement**
- 4) $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ **Synthesis**
- *5) $\text{Ca}(\text{OH})_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$ **Double Replacement**
- 6) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ **Decomposition**
- 7) $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$ **Single Replacement**
- 8) $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$ **Double Replacement**
- 9) $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$ **Synthesis**
- *10) $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ **Single Replacement**

*Remember: "H₂O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 11) Double replacement: $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 12) Single replacement: $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 13) Decomposition: $\text{Mg}(\text{ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 14) Synthesis: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 15) Double replacement: $3\text{Ca}(\text{OH})_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe}(\text{OH})_3 + 3\text{CaCl}_2$
- 16) Single replacement: $\text{Fe} + \text{Cu}(\text{NO}_3)_2 \rightarrow \text{Fe}(\text{NO}_3)_2 + \text{Cu}$ [Assume Fe²⁺]
- 17) Decomposition: $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 18) Synthesis: $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 19) Double replacement: $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 20) Single replacement: $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$ [Copper (II) is used here]

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Becker, Benjamin L. Carroll

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Paul A. Bartlett, Judith G. Koch, 1992 **Organic Chemistry Study Guide** Robert J. Ouellette, J. David Rawn, 2014-11-04
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Protein-Ligand Interactions Holger Gohlke, 2012-04-06 Innovative and forward looking this volume focuses on recent achievements in this rapidly progressing field and looks at future potential for development. The first part provides a basic understanding of the factors governing protein-ligand interactions followed by a comparison of key experimental methods calorimetry, surface plasmon resonance, NMR used in generating interaction data. The second half of the book is devoted to *in silico* methods of modeling and predicting molecular recognition and binding, ranging from first principles based to approximate ones. Here as elsewhere in the book emphasis is placed on novel approaches and recent improvements to established methods. The final part looks at unresolved challenges and the strategies to address them. With the content relevant for all drug classes and therapeutic fields this is an inspiring and often consulted guide to the complexity of protein-ligand interaction modeling and analysis for both novices and experts.

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Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter, 1993

Study Guide with Selected Solutions Karen

Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems

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Maryland School Bulletin Maryland. State Department of Education, 1927

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Chemistry, Student Study Guide John A. Olmsted, Gregory M.

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The

Journal of Education, 1911

Fundamentals of Chemistry, Study Guide James E. Brady, John R. Holom, 1988-04-20

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descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises

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Study guide to accompany Drew H. Wolfe: General, organic and biological chemistry Daniel C. Pantaleo,1986

Study Guide with Solutions for Organic Chemistry, Third Edition Ralph J. Fessenden,Joan S. Fessenden,1986

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General, Organic, and Biological Chemistry Study Guide and Selected Solutions Karen C. Timberlake,2001-11

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