

CHAPTER 1

MATTER AND MEASUREMENT

Chemistry: deals with composition, structure, and reactions of matter

- **Matter** – has mass and occupies space
 - Mass – measures quantity of matter
 - Weight – result of gravitational attraction between matter
 - Weight and mass are NOT the same
- **Composition** – what matter is made of and how much of each component is present
 - Expressed by...
 - Weight/mass
 - Volume
 - Percent
 - # of moles
 - # of atoms
 - *Macroscopic* – can be seen and weighed
 - *Submicroscopic level* – number and types of atoms
 - *Atoms* = simple units of matter
 - *Molecules* = combinations of atoms
 - Don't separate when dissolved like compounds → ionic compounds don't exist as molecules and separate when dissolved
 - *Qualitative* = ethanol consists of carbon, hydrogen and oxygen
 - *Quantitative* = ethanol = C₂H₆O
- **Structure** – arrangement of components and how they are bonded
 - Different arrangements of same combination of components = isomers
 - Structures rotate freely
 - Dashed lines = behind dark triangles = forward
- **Reactions** – changes in composition and structure
 - What products?
 - How much of product? (stoich)
 - How fast change occurs?
 - What energy changes accompany rxn?



- exothermic reaction → produces energy over a long period of time

Ethanol
C₂H₆O
or
C₂H₅OH



Scientific Method

- Experiment (record observations) → numerical + observational
 - Careful recordings and analysis of data
 - Controlled conditions so we pinpoint what is causing another thing to happen
 - Reproducible
- Draw a Conclusion – **LAW**
 - Concise statement about a basic relationship or regularity of nature drawn from observations

Chemistry Chapter 1 Study Guide

Darrell D. Ebbing, Susan M. Schelble



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