

Chemistry | Semester 2 | Final Exam Study Guide
Unit 9: Compound Stoichiometry

38. A mole is a unit defined to be $= 6.02 \times 10^{23}$ particles
1 mole = 1 mol

39. The molar mass of an element is the mass of one mole of the substance
= 19.5 atomic mass in grams

40. What information is given to you in the empirical formula?
- elements in comp.
- ratio of elements

41. What do the subscripts represent in a molecular formula?
how many atoms of each element

42. Define Formula Mass
mass of 1 mole of an ionic compound

43. Define Molar Mass
mass of 1 mole of any compound

44. Define Molecular Mass
mass of one mole of a covalent (molecular) comp.

True/False

32. F One mole of O₂ has a greater mass than 1 mole of Na₂O.

33. F One mole of H₂ has more atoms than 1 mole of Na₂O. (shorter volume at same temp)

34. F One molecule of H₂ has more atoms than 1 molecule of SO₂. (3 atoms)

Calculations

45. Find the molar mass of Al₂(SO₄)₃.

Find mass of each element & multiply each mass by subscript & add masses of all elements together = molar mass

$$\text{Al } (2) 26.98 = 53.96 \\ \text{S } (3) 32.07 = 96.21 \\ \text{O } (12) 16.00 = 192.00 \\ \text{Total: } 294.17 \text{ g/mol}$$

46. How many moles are in 2.4g of carbon dioxide (CO₂)?

Start with the given value & use dimensional analysis or you can factor out labels until you get to target units

$$2.4 \text{ g CO}_2 | 1 \text{ mol CO}_2 = 0.05 \text{ mol} \\ 44 \text{ g/mol}$$

molar mass CO₂

47. How many molecules are contained in 35.37g of water? (molar mass of water is 18.02g/mol)?

$$35.37 \text{ g H}_2\text{O} | 1 \text{ mol H}_2\text{O} | 6.02 \times 10^{23} \text{ particles} \\ 18.02 \text{ g H}_2\text{O} | 1 \text{ mol H}_2\text{O}$$

$$1.18 \times 10^{24} \text{ molecules H}_2\text{O}$$

51. How many atoms are in 6.3 moles of calcium?
Note: There are 6.02×10^{23} atoms in 1 mole

$$6.3 \text{ mol Ca} | 6.02 \times 10^{23} \text{ atoms} = 3.79 \times 10^{24} \text{ atoms} \\ 1 \text{ mol Ca}$$

52. Calculate the number of moles there are in 4.5×10^{24} molecules of sulfur dioxide (SO₂).

$$4.5 \times 10^{24} \text{ molecules SO}_2 | 1 \text{ mol SO}_2 = 747.51 \text{ mol SO}_2 \\ 6.02 \times 10^{23} \text{ particles}$$

53. Calculate the percent composition of sodium and oxygen in sodium oxide (Na₂O).

Find the molar mass → divide total mass of element within Na₂O by the total mass of compound

$$\text{Na } (2) 22.99 = 45.98 / 61.98 = 74.2\% \\ \text{O } (2) 16.00 = 32.00 / 61.98 = 25.8\%$$

molar mass

$$\frac{45.98}{61.98} \text{ g/mol}$$

54. What is the empirical formula of a compound that contains 81.68% carbon and 18.32% hydrogen?

assume 100 g of % to mass → mass to mole → divide by smallest → multiply all whole

$$8.68 \text{ g C} | 1 \text{ mol} = 6.8 / 6.8 = 1 \times 3 = 3 \\ 12.01 \text{ g C}$$

$$18.32 \text{ g H} | 1 \text{ mol H} = 18.139 / 6.8 = 2.67 \times 3 = \\ 1.01 \text{ g H}$$

$$\text{C}_3\text{H}_8$$

55. What is the molecular formula of a hydrocarbon that has an empirical formula of CH₄ and a molecular mass (molar mass of molecule) of 48.15 g/mol?

$$\frac{\text{MM}_{\text{mole}}}{\text{MM}_{\text{emp}}} = \frac{48.15}{16.05} = 3 \text{ C}_3\text{H}_4$$

Multiply each subscript by the above answer to get the subscript for the molecular formula

To find Empirical Formula
#54 (multiply till whole)

if # ends in 1 multiply by

0.2 or 0.8	× 5
0.25 or 0.75	× 4
0.33 or 0.67	× 3
0.5	× 2

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