

Good Friday Worksheet

Period 4.8.8: Classification (Homework, 10 marks)

1. Identify which type of bond has the following properties.
 (Answer: Ionic, Metallic, Nonpolar, Covalent)

Metallic bond	Ionic bond	Covalent bond
<ul style="list-style-type: none"> Size of delocalized electrons Conduction when solid 	<ul style="list-style-type: none"> Oppositive charges Large resulting ionic radius Electrostatic attraction between delocalized electrons 	<ul style="list-style-type: none"> Clear ion solid, liquids, or gases Smaller resulting ionic radius Stronger ionic bond

2. Circle the correct word that will complete the sentence below.

Nonpolar bonds result from ~~small~~ electronegativity difference. The element with a stronger electronegativity will ~~attract~~ electrons from the element with a weaker electronegativity. Covalent bonds result in ~~large~~ electronegativity difference. So, the elements will ~~share~~ their electrons with each other in a ~~nonpolar~~ ~~nonpolar~~ way.

3. Fill in the chart below using the element names or element symbols.

Symbol or Element	Formula	Name	Lewis Structure	pn Covalency	pn Covalency Shape
Tin(II)	Sn_{2+}	Stannous Ions	[Sn]^{2+}	Nonpolar	Nonpolar
Carbon	C_{60}	Carbon nanotubes	[C]_{60}^{2-}	Nonpolar	Tetrahedral
Tin(IV)	SnO_4	Copper(II) nitrate	$\text{[Sn]}^{4+} \text{[O]}_4^{2-}$	Polar	Tetrahedral
Carbon	CO_2	Oxygen difluoride	$\text{[O]}_2 \text{[F]}_2$	Polar	Linear
Tin(II)	Sn_{2+}	Stannous Iodide	$\text{[Sn]}^{2+} \text{[I]}^-$	Nonpolar	Linear
Tin(IV)	SnO_2	Copper(II) sulfide	$\text{[Sn]}^{4+} \text{[O]}_2 \text{[S]}^{2-}$	Polar	Tetrahedral
Carbon	C_{60}	Carbon triiodide	$\text{[C]}_{60}^{2-} \text{[I]}_3$	Nonpolar	Trigonal planar
Carbon	H_2S	Hydrogen sulfide	$\text{H} - \text{S} - \text{H}$	Polar	Linear

Short-Block Questions

Important Vocabulary & their Definition:

- Wavelength: ~~long~~ ~~it~~ ~~is~~ ~~the~~ ~~size~~ ~~of~~ ~~one~~ ~~wave~~ ~~in~~ ~~the~~ ~~medium~~
- Frequency: ~~small~~ ~~it~~ ~~is~~ ~~the~~ ~~size~~ ~~of~~ ~~one~~ ~~wave~~

1. Circle the correct word that will complete the sentence below.

Elements with a large electronegativity will have a ~~large~~ frequency and ~~large~~ energy.

2. Calculate the energy (E) and wavelength (λ) of a photon of light with a frequency (f) of $5.165 \times 10^{14} \text{ Hz}$.

$$E = hf$$

$$= (6.626 \times 10^{-34} \text{ J})(5.165 \times 10^{14} \text{ Hz})$$

$$E = 3.39 \times 10^{-19} \text{ J} = \frac{3.39}{1.602 \times 10^{-19} \text{ J}}$$

$$\lambda = \frac{c}{f} = \frac{3.0 \times 10^8 \text{ m/s}}{5.165 \times 10^{14} \text{ Hz}} = 5.8 \times 10^{-7} \text{ m}$$

$$E = 6.626 \times 10^{-34} \text{ J} \times 5.165 \times 10^{14} \text{ Hz}$$

$$E = 3.39 \times 10^{-19} \text{ J}$$

Chemistry Final Study Guide Answer Key

Margaret G. Kimble

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