

Notes:**Properties of solids, liquids, & gases:**

Solids	Liquids	Gases
<ul style="list-style-type: none"> ○ Particles very close together ○ Highest density ○ Lowest kinetic energy ○ Least amount of motion ○ Definite shape & volume 	<ul style="list-style-type: none"> ○ Particles loosely held together ○ Moderate density ○ Moderate kinetic energy ○ Some motion ○ No definite shape, definite volume 	<ul style="list-style-type: none"> ○ Particles very far apart ○ Least dense ○ Most kinetic energy ○ High amount of motion ○ No definite shape or volume ○ No organization

Phase Changes:Solid \rightarrow liquid: meltingLiquid \rightarrow solid: freezinggas \rightarrow liquid: condensationLiquid \rightarrow gas: evaporating/vaporizinggas \rightarrow solid: depositionsolid \rightarrow gas: sublimation**Gases:**

There are three main characteristics of gases – all of which are a result of the kinetic molecular theory which says:

1. All particles of a gas are in constant motion
2. As a result of this motion, random collisions occur, but gases are not attracted or repelled from each other
3. Gases are much smaller than the distance between them.

The characteristics of gases are:

1. Volume: the amount of space a gas occupies – usually measured in liters (L) or milliliters (mL)
2. Pressure: force per unit area – comes from particles banging into the side of the container. Measured in kilopascals, atmospheres, or mm of mercury.
3. Temperature: average kinetic energy of the particles of gas – measured in degrees Celsius or Kelvin.
4. STP: standard temperature & pressure

Pressure: 1 atm = 101.3 kPa = 760 mmHg = 760 torr

Temperature: 0°C = 273 K K = °C + 273

Volume: 1 mole = 22.4 L

Scientific Laws & Discoveries:

1. Dalton's Law of Partial Pressure: the total pressure of a gas is equal to the sum of the pressures of each part of the gas mixture.
 $P_{\text{total}} = P_1 + P_2 + P_3 + \dots + P_n$
2. Graham's Law of Effusion: at constant temperature and pressure, the rate of effusion is inversely proportional to the square root of their molar masses. $\text{rate of effusion}_{\text{gas A}} / \sqrt{\text{molar mass}_{\text{gas A}}} = \text{rate of effusion}_{\text{gas B}} / \sqrt{\text{molar mass}_{\text{gas B}}}$
3. Boyle's Law: if the temperature remains the same, pressure & volume are indirectly proportional. If one condition goes up, the other goes down & vice versa. $P_1V_1 = P_2V_2$
4. Charles's Law: if the pressure remains constant, the temperature & volume are directly proportional. If one condition goes up, so does the other. $\frac{V_1}{T_1} = \frac{V_2}{T_2}$
5. Gay-Lussac's Law: if the volume remains constant, the pressure & temperature are directly proportional – if one condition goes up, so does the other. $\frac{P_1}{T_1} = \frac{P_2}{T_2}$
6. **Combined Gas Law:** shows the previous 3 gas laws in one. To use any of the 3 previous laws, simply eliminate the variable that remains constant. $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
7. **Ideal Gas Law:** Avogadro came along and said that perhaps the number of particles (atoms or molecules) in the gas sample may affect the three characteristics of gases and he was right. So, the number of moles (n) in the sample also has an effect. You may have to convert from grams to moles, moles to molecules, or vice versa depending on what the question gives/wants.

$PV = nRT$ R is the gas law constant (there are 3 different ones you can use based on the unit the pressure is given in).
 $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$ $8.31 \frac{\text{J}}{\text{mol} \cdot \text{K}}$ $62.4 \frac{\text{L} \cdot \text{mmHg}}{\text{mol} \cdot \text{K}}$

**Be careful to make sure the units cancel!!!

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Rachel S Tattersall

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