

Gas Laws Study Guide

Test Format:

- Multiple choice and problems

Items that will be provided for the test (read important notes at the bottom):

- Gas law equations
- The value of R
- Pressure conversion factors
- I will provide the equations on the test. You are responsible for knowing which laws the equations describe (ex: I will give you $P_1 \times V_1 = P_2 \times V_2$ and you must know that this is Boyle's Law).

- What is the name of the model we currently use to explain the behavior of gases? kinetic molecular theory (KMT).
- Define the kinetic molecular theory (KMT). **Gases are in constant motion with elastic collisions.**
- The KMT makes many assumptions about the **size**, **motion**, and **energy** of gas particles.
- According to the KMT, do gas particles experience attractive or repulsive forces? Explain why or why not. **Neither- they are too far apart**
- Completely describe the motion of gas particles according to the KMT. **Constant, Random, Straight paths, different velocity**
- Define elastic collision. **No kinetic energy is lost- just transferred between particles**
- What determines the kinetic energy of a particle? **Mass and velocity**
- In a sample of a single gas, do all particles have the same kinetic energy? Explain why or why not. **No – they have a different velocity**
- Define temperature. **Measurement of kinetic energy**
- Explain why gases have a low density. **Lots of space between particles (large volume)**
- Why are gases compressible? **Lots of empty space between particles**
- Why do gases expand to fill a container? **No attractive or repulsive forces**
- Why can gas particles flow easily past each other? **No attractive forces**
- Define diffusion. **Movement of one material through another**
- Define effusion. **Gas escapes through tiny openings**
- Describe how diffusion depends on the mass of the particles. **Diffusion is dependent on mass, so less mass more diffusion**
- Use the KMT to explain why a balloon expands as you blow air into it. **Particles are in-constant motion and collide with their surroundings expanding the balloon**

Boyle's Law: Answer the following questions about Boyle's Law.

- Describe Boyle's Law in your own words. **As pressure increases, volume decreases**
- What is the formula? $P_1V_1 = P_2V_2$
- What are the variables? What are the possible units used in the formula? **Pressure and volume, atm, kPa, mL, L**
- What is constant? **temperature**



- Give an example to help explain the law (you should have some type of visual of your example).
- What would a graph of pressure vs. volume of a gas look like according to Boyle's Law?

Charles' Law: Answer the following questions about Charles' Law.

- Describe Charles' Law in your own words. **Increase temperature, increase volume**
- What is the formula? **=**
- What are the variables? What are the possible units used in the formula? **Temperature and volume (kelvin, mL, L)**
- What is constant? **Pressure**
- What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- Give an example to help explain Charles's Law (you should have some type of visual of your example).



- What would a graph of temperature vs. volume of a gas look like according to Charles's Law?

Gay-Lussac's Law: Answer the following questions about Gay-Lussac's Law.

- Describe Gay-Lussac's Law in your own words. **As pressure increases, temperature increases**
- What is the formula? **=**
- What are the variables? What are the possible units used in the formula? **Pressure and temperature (kelvin, atm, kPa, torr)**
- What is constant? **volume**
- What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- Give an example to help explain Gay-Lussac's Law (you should have some type of visual of your example).

Chemistry Gases Study Guide

G. Marc Loudon, Joseph G. Stowell

Chemistry Gases Study Guide:

Safety Professional's Reference and Study Guide W. David Yates,2017-12-12 While there are numerous technical resources available often you have to search through a plethora of them to find the information you use on a daily basis And maintaining a library suitable for a comprehensive practice can become quite costly The new edition of a bestseller Safety Professional's Reference and Study Guide Second Edition provides a single source reference that contains all the information required to handle the day to day tasks of a practicing industrial hygienist New Chapters in the Second Edition cover Behavior based safety programs Safety auditing procedures and techniques Environmental management Measuring health and safety performance OSHA's laboratory safety standard Process safety management standard BCSPs Code of Ethics The book provides a quick desk reference as well as a resource for preparations for the Associate Safety Professional ASP Certified Safety Professional CSP Occupational Health and Safety Technologist OHST and the Construction Health and Safety Technologist CHST examinations A collection of information drawn from textbooks journals and the author's more than 25 years of experience the reference provides as the title implies not just a study guide but a reference that has staying power on your library shelf

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Safety Professional's Reference and Study Guide, Third Edition W. David Yates,2020-03-19 This new edition serves both as a reference guide for the experienced professional and as a preparation source for those desiring certifications It's an invaluable resource and a must have addition to every safety professional's library Safety Professional's Reference and Study Guide Third Edition is written to serve as a useful reference tool for the experienced practicing safety professional as well as a study guide for university students and those preparing for the Certified Safety Professional examination It addresses major topics of the safety and health profession and includes the latest version of the Board of Certified Safety Professional BCSP reference sheet a directory of resources and associations as well as state and federal agency contact information Additionally this new edition offers new chapters and resources that will delight every reader This book aids the prospective examination candidate and the practicing safety professional by showing them step by step how to solve each question formula listed on the BCSP examination and provide examples on how and when to utilize them

Study Guide to Accompany Calculus for the Management, Life, and Social Sciences Clyde Metz,1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences

Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl,Paul B. Kelter,1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

Self Study Guide B. Pharma

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TOC Solved Papers 2020 2012 Physics Chemistry Botany Zoology Appendix [Chemical Principles Student's Study Guide & Solutions Manual](#) John Krenos,Joseph Potenza,Loretta Jones,Lynn Koplitz,Thomas Spence,2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter,1993 **Chemical Principles Study Guide/Solutions Manual**

John Krenos,Joseph Potenza,Carl Hoeger,2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

[Inorganic Chemistry Fast Facts: Gas Phase](#) E Staff, Learn and review on the go Use Quick Review Chemistry Study Notes to help you learn or brush up on the subject quickly You can use the review notes as a reference to understand the subject better and improve your grades Easy to remember facts to help you perform better Perfect study notes for all high school college health sciences premed medical and nursing students [Chemistry - Study Guide to Accompany](#)

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State Manual of the Courses of Study for the High Schools of Oregon ... Oregon. Office of Superintendent of Public Instruction,1925 **Solutions Manual and Study Guide to Accompany Introduction to Organic Chemistry, 4th Ed** Paul A. Bartlett,Judith G. Koch,1992

[Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition](#) Susan McMurry,2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters

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