

## Chemistry Study Guide: Unit 6: Gases

**Kinetic Molecular Theory (KMT):** a model or theory that is used to explain the behavior of gases.

- A. Gases contain particles (molecules or atoms) that are in constant, random, straight-line motion.
- B. Gas particles collide with each other and with the walls of the container. These collisions are perfectly elastic, meaning kinetic energy (energy an object has because of its motion) is being transferred among the particles but no kinetic energy is lost.
- C. Gas particles are separated by relatively great distances. Because of this, the volume of the particles themselves is negligible (insignificant).
- D. Gas particles do not attract or repel each other. This is because they are far away from each other.

**Gas Pressure:** Force exerted by a gas on the walls of its container.

- Caused by collisions between gas particles and walls of the container.
- Units: kilopascal (kPa), Atmosphere (atm), millimeters mercury (mm Hg)
- $1 \text{ atm} = 760 \text{ mmHg} = 101.3 \text{ kPa}$
- Found on table A with temperature on periodic table

**Temperature:** a measure of the average kinetic energy of the particles in a substance.

- As average kinetic energy of particles in a substance increases, temperature of a substance increases.
- In science, Celsius and kelvin are common temperature scales.
- $K = C^{\circ} + 273$ ,  $C^{\circ} = K - 273$
- Temperature is not a form of energy.

Relationship between pressure and volume of a gas:

**Boyle's Law:** Pressure and Volume (Temperature Constant)

If the temperature is constant, then the pressure of a gas increases, as the volume decreases. As the pressure decreases the volume increases.

- Inverse relationship between pressure and volume.

**Equation:**  $P_1 \bullet V_1 = P_2 \bullet V_2$

**Charles Law:** Temperature and Volume (Pressure Constant)

If the pressure is constant, then as the temperature of an enclosed gas increases, the volume increases. As the temperature decreases, the volume decreases.

- Direct relationship between temperature and volume
- Temperature must be converted to Kelvin.

**Equation:**  $V_1 / T_1 = V_2 / T_2$

**Gay-Lussac's Law:** Temperature and Pressure (Volume is constant)

If the volume is constant, then as the temperature of an enclosed gas increases, the pressure increases. As the temperature decreases, the pressure decreases.

- Direct relationship between temperature and pressure.
- Temperature must be converted to Kelvin.

**Equation:**  $P_1 / T_1 = P_2 / T_2$

# Chemistry Gases Unit Study Guide

**John S. Phillips, Cheryl Wistrom**



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**Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Study Guide [to] Chemistry, a Systematic Approach [by] Sisler, Dresdner, and Mooney* William H. Myers, William T. Mooney, 1980 **Resources in education**, 1984-06 *Study Guide* Darrell D. Ebbing, Susan M. Schelble, 1998 **Gage Canadian GED Study Guide Series** Kathy Austin, 1998

Chemistry - Study Guide to Accompany Moore/Davies/Collins Clare T. Furse, 1978 **McGraw-Hill Encyclopedia of Science and Technology** Daniel N. Lapedes, 1977 Accompanied by Readers guide 33 p 28 cm New York McGraw Hill c1977 Call number Q121 M3 1977 Suppl 1 Accompanied by Study guide 50 p 28 cm New York McGraw Hill c1977 Call number Q121 M3 1977 Suppl 2 **Gas Chromatography** Ian A. Fowles, 1995-06-14 This volume has been revised and updated to include new chapters on high resolution gas chromatography GC GC detectors and the classification of GC Self assessment questions are used throughout the text to ensure a full understanding of unfamiliar concepts El-Hi Textbooks in Print, 1980 **Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan McMurry, 2000

Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

**McGraw-Hill encyclopedia of science & technology** McGraw-Hill, 2002 Barron's how to Prepare for College Entrance Examinations Samuel C. Brownstein, Mitchel Weiner, 1974 A guide to preparing for college entrance examinations with emphasis on study programs for the verbal mathematics and standard written English parts of the SAT Includes practice tests

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**Chemistry** John S. Phillips, Cheryl Wistrom, 2000 Gas Journal, 1930 **Resources in Education**, 1982 *Medical Technology Examination Review and Study Guide* Anna P. Ciulla, Georganne K. Buescher, Joan H. Youse, 1988 *The Indiana Teacher*, 1936 Iron & Coal Trades Review, 1929 Vol 115 includes Diamond jubilee issue 1867 1927 **Study Guide for Chemical Principles [by]**

**Steven S. Zumdahl** Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach

taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

## **Chemistry Gases Unit Study Guide** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Chemistry Gases Unit Study Guide**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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