

Gases Test SCH3U

Name:

K 15 T 9 C 5 A 10 **Assessment of Learning (Product)**

Multiple Choice. Circle the most correct answer. (15K)

1. Convert 120 K to °C.

A) 393 °C B) -153 °C C) 880 °C D) 153 °C

2. Why do we assume that gas particles experience no intermolecular forces?

A) Because they are moving too quickly B) Because they are small.
C) Because gas particles are a long distance from each other D) None of the above

3. The mass of a fixed quantity of a gas:

A) increases when the volume increases, at constant temperature
B) decreases when the absolute temperature decreases, at constant temperature
C) does not change at any temperature and pressure
D) decreases as the density increases
E) decreases when the volume decreases, at constant pressure

4. Which rigid cylinder contains the same number of gas molecules at STP as a 2.0- liter rigid cylinder containing H₂ (g) at STP?

A) 2.0- L cylinder of CH₄ (g)
B) 1.5- L cylinder of NH₃ (g)
C) 4.0- L cylinder of He (g)
D) 1.0- L cylinder of O₂ (g)

5. Which statement describes the particles of an ideal gas based on the kinetic molecular theory?

A) The gas particles are relatively far apart and have negligible volume.
B) The gas particles are in constant, nonlinear motion.
C) The gas particles have attractive forces between them.
D) When gas particles have collisions they lose energy.



6. Which phase depicted above does not have a definite shape or volume?

A) The one in the middle B) The one on the left
C) The one in the middle and the one on the right D) The one on the right

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