

## Chemistry - Chapter 7 Study Guide

### Section 7.1:

- *Valence electrons* – electrons in the outer energy level
  - responsible for bonding with other atoms
- How do you know how many valence electrons an element has?
  - Look at the group number
    - Group 1A = 1 valence electron
    - Group 7A = 7 valence electrons
    - Etc...
- *Electron Dot Diagrams* – a way of showing and keeping track of valence electrons
  - Written using the following steps:
    1. Write the element symbol – representing the nucleus and inner electrons
    2. Put one dot for each valence electrons.
    3. Don't forget to add them using Hund's rule – they don't pair up unless they have to.
- *The Octet Rule* – atoms tend to want to achieve a noble gas configuration because it is the most stable. Octet = eight electrons
- *Cations* – a positive ion (usually metals)
  - Lose electrons to attain a noble gas configuration
  - Written like  $\text{Na}^{1+}$
  - No dots are shown (on an electron dot diagram) because they lose the valence electrons they had
- *Anions* – a negative ion (usually non-metals)
  - Gain electrons to attain a noble gas configuration
  - Written like  $\text{P}^{3-}$
  - Dots are shown because valence electrons are gained
- 8 valence electrons = stable atom (noble gases.)

### Section 7.2:

- Anions (-) and cations (+) are held together by opposite charges; a bond is formed through the transfer of electrons to achieve a noble gas configuration (octet.)
- *Ionic Compounds:*
  - Also called salts
  - Made from a cation and an anion
    - Which means they're made of a metal and a non-metal
- **"if you don't understand how to do ionic bonding, check the book or ask somebody to help... it's hard to explain"**
- The result of ionic bonding is an ionic compound, shown in a *formula unit*
  - Written like  $\text{Ca}_3\text{P}_2$

# Chemistry Matter And Change Chapter 7 Study Guide Answer Key

**DJ Losen**



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### **Table of Contents Chemistry Matter And Change Chapter 7 Study Guide Answer Key**

1. Understanding the eBook Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - The Rise of Digital Reading Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Matter And Change Chapter 7 Study Guide Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Matter And Change Chapter 7 Study Guide Answer Key

- Personalized Recommendations
  - Chemistry Matter And Change Chapter 7 Study Guide Answer Key User Reviews and Ratings
  - Chemistry Matter And Change Chapter 7 Study Guide Answer Key and Bestseller Lists
5. Accessing Chemistry Matter And Change Chapter 7 Study Guide Answer Key Free and Paid eBooks
    - Chemistry Matter And Change Chapter 7 Study Guide Answer Key Public Domain eBooks
    - Chemistry Matter And Change Chapter 7 Study Guide Answer Key eBook Subscription Services
    - Chemistry Matter And Change Chapter 7 Study Guide Answer Key Budget-Friendly Options
  6. Navigating Chemistry Matter And Change Chapter 7 Study Guide Answer Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Chemistry Matter And Change Chapter 7 Study Guide Answer Key Compatibility with Devices
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