

CHEMISTRY CHEAT SHEET

INTERMOLECULAR FORCES

INTERMOLECULAR FORCES CHEAT SHEET

Intermolecular forces are the forces of attraction or repulsion between molecules. These forces are responsible for the physical properties of matter.

The stronger the intermolecular forces, the greater the physical properties of a substance. The stronger the intermolecular forces, the higher the melting point and boiling point.

There are four types of intermolecular forces: ionic, covalent, hydrogen bonding, and van der waals.

Ionic forces are the strongest type of intermolecular force. They are the forces of attraction between ions.



Very strong
Very high
Very high

Covalent forces are the second strongest type of intermolecular force. They are the forces of attraction between atoms.



Very strong
Very high
Very high

Hydrogen bonding is a type of covalent force. It is the attraction between a hydrogen atom and an oxygen atom.



Very strong
Very high
Very high

Van der waals forces are the weakest type of intermolecular force. They are the forces of attraction between molecules.



Very strong
Very high
Very high

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Chemistry Physics Forces Study Guide

G. A. F. VAN RHYN

Chemistry Physics Forces Study Guide:

Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee. Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans. Now that's real chemical understanding. Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist. This text helps you understand how chemical phenomena are governed by what happens at the molecular level, apply critical thinking skills to chemical concepts and problems, and master the basic mathematical techniques needed for quantitative reasoning. You'll see the world as chemists do and learn to appreciate the chemical processes all around us. A Fourth Edition with a lot of new perks. Revisions include a new early energy chapter, revised coverage of bonding, expanded coverage of intermolecular forces, and increased coverage of multiple equilibria including polyprotic acids. New pedagogy strengthens students' critical thinking and problem solving skills. Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills, concepts, equations, and terms. eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text, links between problems and relevant sections in the online text, practice quizzes, the Visual Tutor, Interactive LearningWare problems, and lab demos, as well as homework management and presentation features for instructors.

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