

Chemistry

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(Chapter 1)(Solutions) (Exercises)
(Class 12)

Question 1.26:

If the density of some lake water is 1.25 g mL^{-1} and contains 92 g of Na^+ ions per kg of water, calculate the molality of Na^+ ions in the lake.

Answer 1.26:

Number of moles present in 92 g of Na^+ ions = $\frac{92 \text{ g}}{23 \text{ g mol}^{-1}} = 4 \text{ mol}$

Therefore, molality of Na^+ ions in the lake = $\frac{4 \text{ mol}}{1 \text{ kg}} = 4 \text{ m}$

Question 1.27:

If the solubility product of CuS is 6×10^{-16} , calculate the maximum molarity of CuS in aqueous solution.

Answer 1.27:

Solubility product of CuS , $K_{sp} = 6 \times 10^{-16}$ Let s be the solubility of CuS in mol L^{-1} .



$$K_{sp} = [\text{Cu}^{2+}][\text{S}^{2-}]$$

Now,

$$= s \times s = s^2$$

Then, we have, $K_{sp} = s^2 = 6 \times 10^{-16}$

$$\Rightarrow s = \sqrt{6 \times 10^{-16}} = 2.45 \times 10^{-8} \text{ mol L}^{-1}$$

Hence, the maximum molarity of CuS in an aqueous solution is $2.45 \times 10^{-8} \text{ mol L}^{-1}$.

Question 1.28:

Calculate the mass percentage of aspirin ($\text{C}_9\text{H}_8\text{O}_4$) in acetonitrile (CH_3CN) when 6.5 g of $\text{C}_9\text{H}_8\text{O}_4$ is dissolved in 450 g of CH_3CN .

Answer 1.28:

6.5 g of $\text{C}_9\text{H}_8\text{O}_4$ is dissolved in 450 g of CH_3CN .

Then, total mass of the solution = $(6.5 + 450) \text{ g} = 456.5 \text{ g}$

Therefore, mass percentage of $\text{C}_9\text{H}_8\text{O}_4 = \frac{6.5}{456.5} \times 100 \% = 1.424 \%$

Question 1.29:

Nalorphine ($\text{C}_{19}\text{H}_{21}\text{NO}_3$), similar to morphine, is used to combat withdrawal symptoms in narcotic users. Dose of nalorphine generally given is 1.5 mg. Calculate the mass of $1.5 \times 10^{-3} \text{ m}$ aqueous solution required for the above dose.

Answer 1.29:

The molar mass of Nalorphine ($\text{C}_{19}\text{H}_{21}\text{NO}_3$) is given as:

$$19 \times 12 + 21 \times 1 + 1 \times 14 + 3 \times 16 = 311 \text{ g mol}^{-1}$$

In $1.5 \times 10^{-3} \text{ m}$ aqueous solution of nalorphine, 1 kg (1000 g) of water contains $1.5 \times 10^{-3} \text{ mol}$

$$= 1.5 \times 10^{-3} \times 311 \text{ g} = 0.4665 \text{ g}$$

Therefore, total mass of the solution = $(1000 + 0.4665) \text{ g} = 1000.4665 \text{ g}$

This implies that the mass of the solution containing 0.4665 g of nalorphine is 1000.4665 g.

Therefore, mass of the solution containing 1.5 mg of nalorphine is: $\frac{1000.4665 \times 1.5 \times 10^{-3}}{0.4665} \text{ g} = 3.22 \text{ g}$

Hence, the mass of aqueous solution required is 3.22 g.

Chemistry Question Form 6

Raffaela Di Napoli

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