



PERIODIC TABLE OF THE ELEMENTS

PERIODIC TABLE OF THE ELEMENTS																		1				
1 H 1.008																		18 Ar 39.948				
3 Li 6.941	4 Be 9.012																	19 K 39.098	20 Ca 40.078	36 Kr 83.80		
11 Na 22.990	12 Mg 24.305																	37 Rb 85.468	38 Sr 87.62	54 Xe 131.29		
19 K 39.098	20 Ca 40.078	21 Sc 44.956	22 Ti 47.88	23 V 50.942	24 Cr 51.996	25 Mn 54.938	26 Fe 55.845	27 Co 58.933	28 Ni 58.693	29 Cu 63.546	30 Zn 65.38	31 Ga 69.723	32 Ge 72.64	33 As 74.922	34 Se 78.96	35 Br 79.904	36 Kr 83.80	55 Cs 132.91	56 Ba 137.33	86 Rn 222		
37 Rb 85.468	38 Sr 87.62	39 Y 88.906	40 Zr 91.224	41 Nb 92.906	42 Mo 95.94	43 Tc 98.906	44 Ru 101.07	45 Rh 102.91	46 Pd 106.42	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.6	53 I 126.91	54 Xe 131.29	55 Cs 132.91	56 Ba 137.33	84 Po 209		
55 Cs 132.91	56 Ba 137.33	57 La 138.905	58 Ce 140.12	59 Pr 140.908	60 Nd 144.24	61 Pm 144.913	62 Sm 150.36	63 Eu 151.964	64 Gd 157.25	65 Tb 158.925	66 Dy 162.50	67 Ho 164.930	68 Er 167.26	69 Tm 168.933	70 Yb 173.054	71 Lu 174.967	85 At 210	86 Rn 222	87 Fr 223	88 Ra 226	118 Uue 294	
87 Fr 223	88 Ra 226	89 Ac 227	90 Th 232.04	91 Pa 231.04	92 U 238.03	93 Np 237.05	94 Pu 244.06	95 Am 243.06	96 Cm 247.07	97 Bk 247.07	98 Cf 251.08	99 Es 252.08	100 Fm 257.10	101 Md 258.10	102 No 259.10	103 Lr 262.11	117 Uut 289	118 Uuq 294	119 Uub 295	120 Uuq 296	116 Uuh 288	
90 Th 232.04	91 Pa 231.04	92 U 238.03	93 Np 237.05	94 Pu 244.06	95 Am 243.06	96 Cm 247.07	97 Bk 247.07	98 Cf 251.08	99 Es 252.08	100 Fm 257.10	101 Md 258.10	102 No 259.10	103 Lr 262.11	104 Rf 261	105 Db 262	106 Sg 266	107 Bh 264	108 Hs 277	109 Mt 268	110 Ds 271	111 Rg 272	112 Uub 285

GENERAL INFORMATION

A. MATTER

- Chemical changes involve the breaking and/or formation of chemical bonds. Physical changes of substance do not result in the formation of new substances.
- In a reaction, matter is neither created nor destroyed, but only changed from one form to another.
- Mass is a measure of the amount of matter in an object.

B. ENERGY

- Energy is the ability to do work or transfer heat.
- Energy appears in such forms as light, sound, heat, or mechanical, electrical, or chemical energy.
- Energy can be converted from one form to another, but can be neither created nor destroyed.
- Two general classifications of energy are potential energy and kinetic energy.

C. ONE MOLE OF ANY SUBSTANCE - is that which contains Avogadro's number (6.02×10^{23}) of particles. Number of moles = mass/molecular mass.

ATOMIC STRUCTURE AND THE PERIODIC TABLE

A. ATOMIC SPECTRA - When an electron moves from the ground state to an excited state, it absorbs energy. When it moves from an excited state to the ground state, it emits energy.

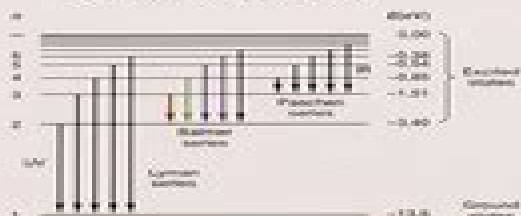
B. THE BOHR THEORY OF THE HYDROGEN ATOM

- Electrons can exist in only certain stable energy levels.
- When the electronic state of an atom changes, it must absorb or emit exactly that amount of energy equal to the difference between the final and initial states.
- The wavelengths of the lines in the spectrum of hydrogen are given by the Rydberg-Ritz formula:

$$\frac{1}{\lambda} = R \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$$

where $R = 109,678 \text{ cm}^{-1}$, n_1 and n_2 are quantum numbers for states a and b , and λ is the wavelength of light emitted or absorbed.

- Transitions between energy levels are shown by the diagram below.

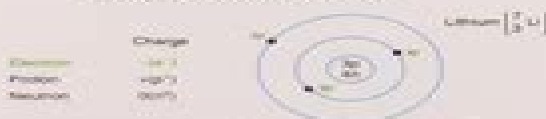


Energy levels of the hydrogen atom with some of the transitions between them that give rise to the spectral lines indicated.

- Light behaves as if it were composed of quanta of energy known as photons: $E_{\text{photon}} = h\nu$, where h is Planck's constant and ν is the frequency of light.

C. ELECTRIC NATURE OF ATOMS

- The nucleus of an atom is composed of positively charged particles called protons and neutral particles called neutrons.
- The atomic number is the number of protons found in the nucleus.
- The number of protons plus the number of neutrons in the nucleus is the mass number of an isotope of an element.
- In a neutral atom, the number of electrons surrounding the nucleus equals the number of protons in the nucleus.
- The electrons found in the outermost shell are called valence electrons.



D. THE WAVE MECHANICAL MODEL

- A wave function corresponds to a certain electronic energy and describes a region about the nucleus (orbital) where an electron having that energy may be found.

- The square of the wave function (Ψ^2), is called a probability density and equals the probability per unit volume of finding the electron in a given region of space.
- An electron can be described with four quantum numbers that can be used to define allowed energy states.
- The four quantum numbers are principle (n), azimuthal (l), magnetic (m), and spin (s).
- Pauli exclusion principle: two electrons within the same atom may not have the same four quantum numbers.

E. ISOTOPES

- Atoms of the same element having different numbers of neutrons are different isotopes of that element.
- Isotopes of an element have virtually the same chemical properties but different radiochemical properties.
- Stability of the nuclei of an isotope of an element is measured by half-life, which is the time it takes for half the atoms of a sample of an isotope to undergo radioactive decay.
- The atomic mass of an element is calculated by averaging the atomic masses of the isotopes of that element.

F. PERIODIC TABLE

- The arrangement of the elements in the periodic table reflects the electronic configurations of the atoms of the elements. Each period in the table corresponds to a principle quantum number and a different outer shell that is filling from left to right.
- Group 18 elements are unreactive because their atoms have full, stable outer shells.
- The periodic law states that chemical and physical properties of the elements are periodic functions of their atomic numbers.
- Vertical columns, called groups, consist of elements that have similar chemical properties because they have the same number of valence electrons.
- The most active metals are in the lower left and the most active nonmetals are in the upper right.
- Metallic properties include high electrical conductivity, high melting points, ductility, and malleability.
- Nonmetals are poor electrical conductors and form brittle solids.
- Metalloids have properties intermediate between those of metals and nonmetals.
- Atomic radii generally decrease across a period from left to right and increase down a group.
- Electronegativity generally increases from left to right in a period and decreases down a group.
- Ionization energy decreases going down a group and increases left to right across a period.

BONDING

A. BONDS

- An ionic bond occurs when one or more electrons are transferred from the valence shell of one atom to another and there is a coulombic attraction between oppositely charged ions. The atom that loses electrons becomes a positive ion (cation), and the atom that acquires electrons becomes a negatively charged ion (anion).
- The octet rule states that atoms tend to gain or lose electrons until their valence shells are filled.
- In a nonpolar covalent bond, the electrons are shared equally between atoms. In a polar covalent bond, there is unequal sharing.

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