



# CHEMISTRY

## CHEAT SHEET

# EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol  $\rightleftharpoons$ .

<p><b>What is Equilibrium?</b></p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.</p>	<p><b>Equilibrium Constant</b></p> <p>The equilibrium constant (<math>K</math>) is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Graph showing concentration vs. time for a reaction at equilibrium.</p> <p>Reactant concentration: <math>[A]_0 - [A]_t = [A]_{eq}</math></p> <p>Product concentration: <math>[C]_0 + [C]_t = [C]_{eq}</math></p>
<p><b>Equilibrium Constant, <math>K_c</math></b></p> <p><math>K_c</math> is the equilibrium constant based on concentrations. It is used for reactions in solution or in the gas phase.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p><b>Equilibrium Constant, <math>K_p</math></b></p> <p><math>K_p</math> is the equilibrium constant based on partial pressures. It is used for reactions involving gases.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	
<p><b>Equilibrium Constant, <math>K</math></b></p> <p><math>K</math> is the general equilibrium constant. It can be used for reactions in solution or in the gas phase.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p><b>Equilibrium Constant, <math>K</math></b></p> <p><math>K</math> is the general equilibrium constant. It can be used for reactions in solution or in the gas phase.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	
<p><b>Le Chatelier's Principle: the effect of changing conditions on equilibrium</b></p>		
<p><b>Concentration</b></p>	<p>Increasing the concentration of a reactant or product shifts the equilibrium in the direction that consumes the added substance.</p>	<p>For the reaction <math>N_2 + 3H_2 \rightleftharpoons 2NH_3</math>, increasing the concentration of <math>N_2</math> shifts the equilibrium to the right, producing more <math>NH_3</math>.</p>
<p><b>Pressure</b></p>	<p>Increasing the pressure shifts the equilibrium in the direction that produces fewer moles of gas.</p>	<p>For the reaction <math>N_2 + 3H_2 \rightleftharpoons 2NH_3</math>, increasing the pressure shifts the equilibrium to the right, producing more <math>NH_3</math>.</p>
<p><b>Temperature</b></p>	<p>Increasing the temperature shifts the equilibrium in the direction that absorbs heat (endothermic reaction).</p>	<p>For the reaction <math>N_2 + 3H_2 \rightleftharpoons 2NH_3</math>, increasing the temperature shifts the equilibrium to the left, producing less <math>NH_3</math>.</p>

# Chemistry Equilibrium Study Guide

**Leo J. Malone**



## Chemistry Equilibrium Study Guide:

**Study Guide for Chemical Principles [by] Steven S. Zumdahl** Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Fundamentals of Chemistry, Study Guide* James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs *Chemistry, Student Study Guide* John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth

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**Resources in Education**, 1984

**Solutions Manual and Study Guide to Accompany Introduction to Organic Chemistry, 4th Ed** Paul A. Bartlett, Judith G. Koch, 1992

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