

**CHAPTER 8****STUDY GUIDE****Covalent Bonding****Section 8.1 The Covalent Bond***In your textbook, read about the nature of covalent bonds.**Use each of the terms below just once to complete the passage.*

covalent bond

molecule

sigma bond

exothermic

pi bond

When sharing of electrons occurs, the attachment between atoms that results is called a(n) (1) \_\_\_\_\_. When such an attachment is formed, bond dissociation energy is released, and the process is (2) \_\_\_\_\_. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) (3) \_\_\_\_\_. If the electrons shared are centered between the two atoms, the attachment is called a(n) (4) \_\_\_\_\_. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) (5) \_\_\_\_\_.

*In your textbook, read about single and multiple bonds and bond strength.**Circle the letter of the choice that best completes the statement or answers the question.*

6. In what form do elements such as hydrogen, nitrogen, and oxygen normally occur?
  - a. as single atoms
  - b. as molecules containing two atoms
  - c. as molecules containing three atoms
  - d. as molecules containing four atoms
7. How many electrons are shared in a double covalent bond?
  - a. none
  - b. one
  - c. two
  - d. four
8. Bond length is the distance between
  - a. two molecules of the same substance.
  - b. the electrons in two attached atoms.
  - c. the nuclei of two attached atoms.
  - d. the orbitals of two attached atoms.
9. Which of the following relationships relating to bond length is generally correct?
  - a. the shorter the bond, the stronger the bond
  - b. the shorter the bond, the weaker the bond
  - c. the shorter the bond, the fewer the electrons in it
  - d. the shorter the bond, the lower the bond dissociation energy

# Covalent Bonding Study Guide Multiple

**James E. Brady, Fred Senese**

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