

## CHARLES' LAW

Name \_\_\_\_\_

Charles' Law states that the volume of a gas varies directly with the Kelvin temperature, assuming that pressure is constant. We use the following formulas:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad \text{or} \quad V_1 \times T_2 = V_2 \times T_1$$
$$K = {}^\circ C + 273$$

Solve the following problems assuming a constant pressure.

1. A sample of nitrogen occupies a volume of 260 mL at 25° C. What volume will it occupy at 95° C? 25 L
2. Oxygen gas is at a temperature of 40° C when it occupies a volume of 2.3 liters. To what temperature should it be raised to occupy a volume of 6.5 liters? 7.5 L
3. Hydrogen gas was cooled from 150° C to 50° C. Its new volume is 75 mL. What was its original volume? 2.5 L
4. Chlorine gas occupies a volume of 26 mL at 300 K. What volume will it occupy at 600 K? 50 L
5. A sample of neon gas at 50° C and a volume of 2.5 liters is cooled to 25° C. What is the new volume? 50 L
6. Fluorine gas at 300 K occupies a volume of 500 mL. To what temperature should it be lowered to bring the volume to 300 mL? 30 L
7. Helium occupies a volume of 3.8 liters at -45° C. What volume will it occupy at 45° C? 38 L
8. A sample of argon gas is cooled and its volume went from 360 mL to 250 mL. If its final temperature was -55° C, what was its original temperature? 25 L

# Charles Law Answer Key With Work

**R Barnett**

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