

# Moles & Stoichiometry Cheat Sheet

<p><b>Calculating Molar Mass:</b></p> <ol style="list-style-type: none"> <li>1. Write out formula of compound</li> <li>2. Determine number of atoms of each element present</li> <li>3. multiply number of atoms of the element by the atomic mass of the element</li> <li>4. add up the values just determined for all of the elements present in the compound</li> </ol>	<p><b>Calculating Percent Composition:</b></p> <ol style="list-style-type: none"> <li>1. Calculate the molar mass of the compound</li> <li>2. divide the mass contribution of each element by the total molar mass</li> <li>3. multiply by 100 to make a percent</li> </ol> <p><i>Note: total of percents should equal 100%!</i></p>
<p>Mass (grams) <math>\xleftarrow{\text{-----}}</math> <math>\xrightarrow{\text{-----}}</math> + by molar mass</p>	<p><math>\times</math> by molar mass <math>\xleftarrow{\text{-----}}</math> Mole <math>\xrightarrow{\text{-----}}</math> <math>\times</math> by Avogadro's # <math>\xleftarrow{\text{-----}}</math> Particles (atoms, molecules, formula units)</p>
<p><b>Calculating Empirical Formula:</b></p> <ol style="list-style-type: none"> <li>1. Change % of each element to grams of each element (if no sample amount is given assume 100 g)</li> <li>2. convert grams of each element to moles of each element by dividing the grams by the molar (atomic) mass of that element</li> <li>3. divide all molar values by the smallest molar value (to establish lowest whole number ratio)</li> <li>4. if all values are not whole numbers, multiply all by a whole number factor to make the values whole (ex. 1.5 and 1 should be multiplied by 2 to get rid of the .5 = 3 and 2)</li> <li>5. plug values in with the element they represent</li> </ol>	<p><b>Calculating Molecular Formula:</b></p> <ol style="list-style-type: none"> <li>1. calculate the empirical formula first</li> <li>2. calculate the molar mass of the empirical formula</li> <li>3. take the molar mass of the molecular formula (given in the problem) and divide it by the molar mass of the empirical formula. This will equal the number of times the e.f. is in the m.f.</li> <li>4. distribute the number attained through the empirical formula</li> </ol> <p><i>Example: <math>2(C_2H_6) = C_4H_{12}</math></i></p>
<p><b>Calculating the Formula of a Hydrate:</b></p> <ol style="list-style-type: none"> <li>1. Determine the mass of water in the hydrate by subtracting the mass of the anhydrous from the mass of the hydrate (if necessary) -OR- convert %'s into grams by assuming a 100g sample</li> <li>2. convert grams of each into moles of each by dividing by molar mass</li> <li>3. divide the moles of water by the moles of the anhydrous to determine the number of moles of water in the formula</li> <li>4. write the formula of the anhydrous followed by a dot then the # of moles of water and "<math>H_2O</math>"</li> </ol> <p><i>Example: <math>CuSO_4 \cdot 5H_2O</math></i></p>	<p><b>Naming a Hydrate:</b></p> <ol style="list-style-type: none"> <li>1. write the name of the ionic compound</li> <li>2. use a prefix for the number of water molecules (see below)</li> <li>3. write "hydrate" after the prefix</li> </ol> <p>1 = mono, 2 = di, 3 = tri, 4 = tetra, 5 = penta, 6 = hexa, 7 = hepta, 8 = octa, 9 = nona, 10 = deca</p> <p><i>Example: Copper (II) sulfate pentahydrate</i></p>

# Cheat Sheet For Stoichiometry

**Jjthetutor**

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was the first volume of the series issued It belongs properly however to this volume and is therefore included here Note p vii  
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Emil J. Margolis,2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction    **Jj's Cheat Sheets** Jjthetutor,2016-04-22 Precalculus Cheat Sheet    **STOICHIOMETRY** NARAYAN CHANDER,2024-04-01 Note Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u@gmail.com You can also get full PDF books in quiz format on our youtube channel <https://www.youtube.com/smartquiziz> I will send you a PDF version of this workbook This book has been designed for candidates preparing for various competitive examinations It contains many objective questions specifically designed for different exams Answer keys are provided at the end of each page It will undoubtedly serve as the best preparation material for aspirants This book is an engaging quiz eBook for all and offers something for everyone This book will satisfy the curiosity of most students while also challenging their trivia skills and introducing them to new information Use this invaluable book to test your subject matter expertise Multiple choice exams are a common assessment method that all prospective candidates must be familiar with in today's academic environment Although the majority of students are accustomed to this MCQ format many are not well versed in it To achieve success in MCQ tests quizzes and trivia challenges one requires test taking techniques and skills in addition to subject knowledge It also provides you with the skills and information you need to achieve a good score in challenging tests or competitive examinations Whether you have studied the subject on your own read for

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