

Calculating the Energy Change

Example: Calculate the enthalpy change when water is formed from H_2 and O_2 .

STEP 1 Bonds Broken

$$2 \times (\text{H}-\text{H}) = 2 \times 436 = 872$$

$$1 \times (\text{O}=\text{O}) = 498$$

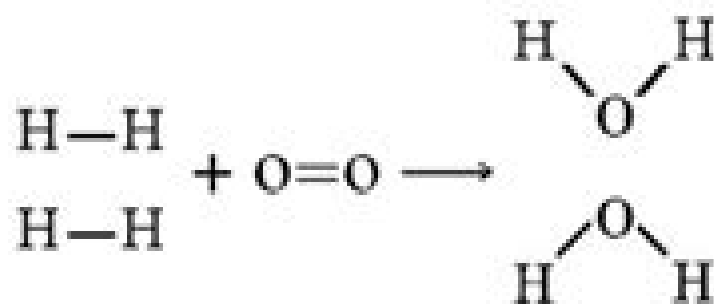
$$\text{Total} = 872 + 498 = 1370$$

STEP 2 Bonds Made

$$4 \times (\text{O}-\text{H}) = 4 \times 464 = 1856$$

STEP 3

$$\begin{aligned}\text{Enthalpy change} &= \text{bonds broken} - \text{bonds made} \\ &= 1370 - 1856 = -486\end{aligned}$$



Bond	Bond Enthalpy
H-H	436
H-O	464
O=O	498

The negative sign means its exothermic.

Calculating Bond Energy

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Calculating Bond Energy:

Chemical Bonds and Bonds Energy R Sanderson, 2012-12-02 Chemical Bonds and Bonds Energy Second Edition provides information pertinent to the fundamental aspects of contributing bond energy and bond dissociation energy This book explores the values that are useful in the interpretation of significant phenomena such as product distribution and reaction mechanisms Organized into 12 chapters this edition begins with an overview of the quantitative relationship among three basic properties of an atom namely nonpolar covalent radius electronegativity and homonuclear single covalent bond energy This text then examines the quantitative means of evaluating the partial atomic charges that result from initial differences in the electromagnetivity of atoms that form a compound Other chapters consider the recognition of the reduction of bond weakening not by multiplicity and in certain types of single covalent bonds The final chapter deals with the application of the principal ideas and techniques to the oxidation of ethane This book is a valuable resource for organic and inorganic chemists

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How Chemical Bonds Form and Chemical Reactions Proceed V. I. Gankin, Yuri V. Gankin, 1998

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