

# CHEMISTRY Terminology

The ABCs of chemistry

## A

**A:** Symbol for atomic mass number  
**absolute zero:** Lowest temperature; 0 Kelvin  
**Ac:** Symbol for the element actinium,  $Z = 89$   
**accuracy:** The closeness of measurement to true value  
**acetate anion:**  $\text{CH}_3\text{COO}^-$ , charge -1  
**acetic acid:**  $\text{C}_2\text{H}_4\text{O}_2$

**acetone:** Dimethyl ketone; organic solvent  
**acid-base indicator:** A chemical that changes color if the pH changes

**acid-base titration:** A method for determining acid or base concentrations  
**acid definitions:** See Arrhenius acid, Brønsted-Lowry acid, Lewis acid

**acid ionization constant ( $K_{\text{a}}$ ):** Equilibrium constant for acid dissociation  
**acidic solution:** pH below 7

**actinides:** Element with  $Z = 50$  or above; radioactive

**activation energy (E<sub>a</sub>):** A process's energy barrier



**affiliation:** Attraction of unlike molecules or materials

**Ag:** Symbol for the element silver,  $Z = 47$

**Al:** Symbol for the element aluminum,  $Z = 13$

**alcohol:** Organic compound with -OH group;  $\text{ROH}$   
**aldehydes:** Carbonyl group bonded to hydrogen and an organic group

**alkali metal family:** Lithium, sodium, potassium, rubidium, cesium, and francium; column #1

**alkaline earth metal family:** Beryllium, magnesium, calcium, strontium, barium, and radium; column #2

**alkane:** Hydrocarbon; all C-C single bonds

**alkene:** Hydrocarbon; 1 or more C=C double bonds

**alkyne:** Hydrocarbon; 1 or more C≡C triple bonds

**allotropes:** Two or more forms of an element

**allyl:** Solution of 2 or more metals

**alpha (α):** Greek letter that denotes radioactive particle and a number of scientific variables

**alpha particle:** Helium nucleus; charge +2

**Am:** Symbol for the element americium,  $Z = 95$

**amalgam:** Alloy of mercury with other metal

**amine:** Organic base;  $\text{R}_3\text{N}$ ; modified ammonia

**amine acid:** Compound with organic acid and organic base property; found proteins and peptides

**ammonium:**  $\text{NH}_4^+$ ; base

**ammonium cation:**  $\text{NH}_4^+$ ; charge +1

**ammonium hydroxide:**  $\text{NH}_4\text{OH}$

**amphoteric oxide:** Exhibits both acid and basic properties

**amu:** Atomic mass unit;  $1/12$  mass of  $\text{C-12}$

**anion:** Ion with a negative charge

**anode:** Electrode that supports oxidation

**antibonding MO:** MO is less stable than separate AOs

**AO:** Abbreviation for atomic orbital

**aq:** Aquous; "dissolved in water"

**Ar:** Symbol for the element argon,  $Z = 18$

**aromatic:** Organic compound with a benzene ring

**Arrhenius acid:** Produces hydronium ion in water solution  
**Arrhenius base:** Yields hydroxide ion in water solution

**Arrhenius equation for rate constant (k):**

$$k = A e^{-E_a/RT}$$

**arsenide anion:**  $\text{As}^{3-}$ ; charge -3

**As:** Symbol for the element arsenic,  $Z = 33$

**At:** Symbol for the element astatine,  $Z = 85$

**atom:** Fundamental unit of all matter

**atomic mass number (A):** Total number of protons and neutrons

**atomic number (Z):** Number of protons in the nucleus

**atomic orbital:** Wave motion of electrons in atoms

**atomic radius:** Empirical measure of atom size

**atomic weight:** Weighted average of natural isotopes of an element

**Au:** Symbol for the element gold,  $Z = 79$

**Aufbau principle:** Guides the filling of electronic subshells of the elements

**Avogadro's law:** Volume is proportional to gas moles at fixed pressure and temperature

**Avogadro's number:**  $N_A = 6.022 \times 10^{23}$  mol<sup>-1</sup>

**B:** Symbol for the element boron,  $Z = 5$

**Ba:** Symbol for the element barium,  $Z = 56$

**balanced equations:** Each side of the equation has the same number of atoms of each element and charge

**balanced redox:** Electron loss = electron gain

**barium cation:**  $\text{Ba}^{2+}$ ; charge +2

**base definition:** See Arrhenius base, Brønsted-Lowry base, Lewis base

**base ionization constant ( $K_b$ ):** Equilibrium constant for base dissociation

**basic solution:** pH over 7

**battery:** Source of electrical power; galvanic cell

**Ber:** Symbol for the element beryllium,  $Z = 4$

**benzene:**  $\text{C}_6\text{H}_6$ ; organic solvent;

aromatic ring of 6 carbon atoms

**beryllium cation:**  $\text{Be}^{2+}$ ; charge +2

**beta (β):** Greek letter that denotes radioactive particle and a number of scientific variables

**beta particle:** Energetic electron from the nucleus; charge -1

**Bh:** Symbol for the element bhabhium,  $Z = 107$

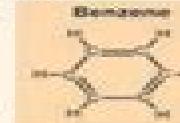
**Bk:** Symbol for the element berkelium,  $Z = 97$

**bicarbonate anion:**  $\text{HCO}_3^-$ ; charge -1

**bimolecular oxygen:**  $\text{O}_2$

**binary compound:** Formed from 2 elements

**bisulfate anion:**  $\text{HSO}_4^-$ ; charge -1



**Bk:** Symbol for the element berkelium,  $Z = 97$

**bleach:** Chemical that is a strong oxidizing agent

**boiling:** Liquid  $\rightarrow$  gas at the boiling point

**boiling point elevation:** A solution has a higher boiling point than a pure solvent

**boiling point (T<sub>b</sub>):** Liquid-gas equilibrium;  $T = 1$  atm

**boiling point of water:** 100°C

**bond energy:** Energy held by a chemical bond

**bond length:** Distance between 2 atoms in a bond

**bond order for Lewis model:** # of bonds divided by # of bonded atoms

**bond order for MO treatment:** # of filled bonding MOs minus # of filled antibonding MOs

**bonding electrons:** Form chemical bonds in a compound

**bonding MO:** MO is more stable than AOs

**Boyle's law:**  $PV =$

constant for gas; Boyle's law

**Br:** Symbol for the element bromine,  $Z = 35$

**bromide anion:**  $\text{Br}^-$ ; charge -1

**brønsted-Lowry acid:** Proton donor

**brønsted-Lowry base:** Proton acceptor

**buffer:** Composed of weak acid and weak base; serves to keep pH constant

## B

**Ba:** Symbol for barium, the SI prefix for "10<sup>3</sup>" or 1000

**c:** Symbol for the speed of light

**°C:** Celsius temperature scale

**C:** Symbol for the element carbon,  $Z = 6$

**C<sub>2</sub>H<sub>4</sub>O<sub>2</sub>:** Acetic acid

**Ca:** Symbol for the element calcium,  $Z = 20$

**calcium cation:**  $\text{Ca}^{2+}$ ; charge +2

**calcium fluoride:**  $\text{CaF}_2$

**calcium hydroxide:**  $\text{Ca}(\text{OH})_2$

**calories:** English unit of heat energy

**calorimetry:** Study of the release or absorption of heat

**carbide anion:**  $\text{C}^-$ ; charge -4

**carboxylate:** Organic compound; carbon bonded to several -COO groups (sugar and starch)

**carbon dioxide:**  $\text{CO}_2$

**carbon monoxide:**  $\text{CO}$

**carbonate anion:**  $\text{CO}_3^{2-}$ ; charge -2

**carbonic acid:**  $\text{H}_2\text{CO}_3$

**carbonyl group:**  $\text{C=O}$ ; ligand complex

**carboxylic group:**  $\text{C}(=\text{O})\text{OOH}$

**catalyst:** Accelerates a reaction but is not consumed in the reaction; lowers the activation energy

**cathode:** Electrode that supports reduction

**cation:** Ion with a positive charge

**Cd:** Symbol for the element cadmium,  $Z = 48$



# College Chemistry 2 Study Guide

**Evan Jay Crane,Austin McDowell  
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