

Name _____ Date _____

CHAPTER 4 STUDY GUIDE FOR CONTENT MASTERY

The Structure of the Atom

Section 4-1 Early Theories of Matter

In your textbook, read about the phlogiston theory, John Dalton, and de Broglie's theory.

For each statement below, write **True** or **False**.**True** _____ 1. Ancient philosophers quickly performed many daily experiments.**True** _____ 2. Philosophers formulated hypotheses about the nature of matter based on their own principles.**True** _____ 3. John Dalton and Robert Boyle proposed that matter is made up of atoms.**True** _____ 4. Dalton's atom theory stated that atoms are equal, uniform, and indestructible in a chemical reaction.**False** _____ 5. Dalton's atom theory stated that matter is made up of empty space.**False** _____ 6. Dalton's atom theory is holding that atoms could not be divided into smaller particles.**False** _____ 7. Dalton's atom theory said that atoms of different elements combine in simple whole number ratios to form compounds.**True** _____ 8. Dalton thought that all atoms of a specific element have the same mass.**False** _____ 9. Democritus proposed his atom as held together by chemical bonds, but no one believed him.**False** _____ 10. Dalton's atom theory was based on visual measurements and inference research.**False** _____ 11. Dalton's atom theory proposed enough empty atoms so that they could be seen.**False** _____ 12. The number of particles of an element that combine to form a particular element is called its atomic mass.

In your textbook, read about the discovery of protons, neutrons, and neutrinos.

Complete the following table of protons, neutrons, and neutrino characteristics.

Particle	Symbol	Location	Relative Charge	Relative Mass
1. Proton	p^+	In the nucleus	1+	1
2. Neutron	n^0	In the nucleus	0	1
3. Electron	e^-	In the space surrounding the nucleus	1-	1/1836

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CHAPTER 4 STUDY GUIDE FOR CONTENT MASTERY

Subatomic Particles and the Nuclear Atom

In your textbook, read about discovering the electron and the nuclear atom.

For each item in Column A, write the letter of the matching item in Column B.

Column A	Column B
1. Proposed the nuclear atom model	a. Thomson
2. Determined the mass-to-charge ratio of an electron	b. Millikan
3. Observed the mass of an electron	c. Rutherford

Draw and label a diagram of each atomic model.

4. Jahn–Teller model

Drawing should look like a ball of chocolate chip cookie dough. The chocolate chip should be labeled with negative charge or electrons. The dough should be labeled as evenly distributed positive charges.

5. nuclear atom model

Drawing should look like a patch with a pit. The pit should be labeled nucleus and should include labeled protons and neutrons. The outer circle of the patch should be labeled electrons.

False _____ 6. Dalton's atom theory proposed enough empty atoms so that they could be seen.**False** _____ 7. Dalton's atom theory was based on visual measurements and inference research.**False** _____ 8. The number of particles of an element that combine to form a particular element is called its atomic mass.

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