

# Stoichiometry and The Mole – Review Guide

## **Remember:**

- 1 mole =  $6.02 \times 10^{23}$  atoms, molecules, formula units
- 1 mole = grams (of molar mass)
- 1 mole = 22.4 liters (L) of gas



1) The process used to predict the amount of a product that will be formed is: **Stoichiometry**

2) If you know the mass of a substance, you can find the number of moles by dividing it by the **molar mass of substance**.

3) If you know the number of particles (atoms) of a substance, you can find the number of moles by dividing it by **Avogadro's Number**.

4) Summarize the steps of a gram-to-gram stoichiometry problem:

Grams A  $\rightarrow$  **Moles A**  $\rightarrow$  **Moles B**  $\rightarrow$  Grams B

5) The essential stoichiometry problems involved in converting from one chemical to another is:

**Mole**  $\rightarrow$  **Mole**

6) Two reactants undergo a chemical reaction. At the end of the reaction, Chemical A still exists in the container, while Chemical B has been completely used up. Which is the limiting reactant?

**Chemical B**

7) What is a mole? **A unit of measurement used to count atoms/molecules.**

8) What is Avogadro's Number?  **$6.02 \times 10^{23}$**

9) What does Avogadro's number tell us? **The number of atoms or molecules in a mole.**

# Chemistry Unit 7 Stoichiometry Review Guide

**John S. Phillips, Cheryl Wistrom**

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