

Gas Laws Study Guide

Lesson 10: Gas Laws	
<p>1. Boyle's Law</p> <p>2. Charles' Law</p> <p>3. Avogadro's Law</p> <p>4. Combined Gas Law</p> <p>5. Gay-Lussac's Law</p> <p>6. Avogadro's Hypothesis</p> <p>7. Ideal Gas Law</p>	<p>8. Gas Laws Study Guide</p> <p>Volume Proportion = $\frac{V_1}{V_2} = \frac{P_2}{P_1}$</p> <p>Volume Proportion = $\frac{V_1}{V_2} = \frac{T_2}{T_1}$ (or $\frac{V_1}{V_2} = \frac{P_1}{P_2} = \frac{T_2}{T_1}$)</p> <p>Volume Proportion = $\frac{V_1}{V_2} = \frac{N_2}{N_1}$</p> <p>Boyle's Law</p> <p></p> <p>$P_1V_1 = P_2V_2$</p> <p>Charles' Law</p> <p>$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$</p> <p>Avogadro's Law</p> <p>$P_1V_1 = P_2V_2$</p> <p>Gay-Lussac's Law</p> <p></p> <p>Compare the following statements to those in the paper. If you are the correct answer, go to the next statement. If not, go to the next question. If you are the correct answer, go to the next statement. If not, go to the next question.</p> <p>1. Pressure increases Temperature increases, because that _____ is correct. (pressure increases, temperature increases)</p> <p>2. Pressure decreases Temperature decreases, because that _____ is correct. (pressure decreases, temperature decreases)</p> <p>3. Pressure increases Temperature decreases, because that _____ is correct. (pressure increases, temperature decreases)</p> <p>4. The number of moles of gas in a container increases the volume increases (because the Pressure decreases (decreases) and the Temperature increases (increases))</p>

Chemistry Study Guide Gas Laws

Steven S. Zumdahl, Paul B. Kelter

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Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg,2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change *Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today* Seager/Slabaugh,2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter,1993 *Chemistry Super Review - 2nd Ed.* Editors of REA,2012-12-15 REA's Chemistry Super Review Get all you need to know with Super Reviews 2nd Edition REA's Chemistry Super Review contains an in depth review that explains everything high school and college students need to know about the subject Written in an easy to read format this study guide is an excellent refresher and helps students grasp the important elements quickly and effectively Our Chemistry Super Review can be used as a companion to high school and college textbooks or as a handy resource for anyone who wants to improve their chemistry skills and needs a fast review of the subject Presented in a straightforward style our review covers the material taught in a beginning level chemistry course including atomic structure bonding chemical reactions liquids solids gases properties of solutions chemical thermodynamics and more The book contains questions and answers to help reinforce what students learned from the review Quizzes on each topic help students increase their knowledge and understanding and target areas where they need extra review and practice **Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry** Iris Stovall,Donald J. DeCoste,2000

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