

**Chemistry - Unit 7 Reaction Equations Worksheet 1**

Balance the following equations by inserting the proper coefficients. For selected reactions, draw *Before* and *After* particle diagrams to show the particles involved in the reaction. Be sure to provide a key.

1. 1 C + 1 H<sub>2</sub>O → 1 CO + 1 H<sub>2</sub>
2. 2 MgO → 2 Mg + 1 O<sub>2</sub>
3. 4 Al + 3 O<sub>2</sub> → 2 Al<sub>2</sub>O<sub>3</sub>

	<b>#3</b>	<b>Before</b>	<b>After</b>
4.	<u>1</u> Zn + <u>1</u> H <sub>2</sub> SO <sub>4</sub> → <u>1</u> ZnSO <sub>4</sub> + <u>1</u> H <sub>2</sub>		
5.	<u>1</u> Cl <sub>2</sub> + <u>2</u> KI → <u>2</u> KCl + <u>1</u> I <sub>2</sub>		
6.	<u>2</u> CuCl → <u>2</u> Cu + <u>1</u> Cl <sub>2</sub>		
7.	<u>2</u> Na + <u>1</u> Cl <sub>2</sub> → <u>2</u> NaCl		
8.	<u>2</u> Al + <u>6</u> HCl → <u>2</u> AlCl <sub>3</sub> + 3 H <sub>2</sub>		

	<b>#8</b>	<b>Before</b>	<b>After</b>
9.	<u>2</u> Fe <sub>2</sub> O <sub>3</sub> → <u>4</u> Fe + <u>3</u> O <sub>2</sub>		
10.	<u>4</u> P + <u>5</u> O <sub>2</sub> → <u>2</u> P <sub>2</sub> O <sub>5</sub>		
11.	<u>1</u> Mg + <u>2</u> HCl → <u>1</u> MgCl <sub>2</sub> + <u>1</u> H <sub>2</sub>		
12.	<u>3</u> H <sub>2</sub> + <u>1</u> N <sub>2</sub> → <u>2</u> NH <sub>3</sub>		
13.	<u>1</u> BaCl <sub>2</sub> + <u>1</u> H <sub>2</sub> SO <sub>4</sub> → <u>1</u> BaSO <sub>4</sub> + <u>2</u> HCl		
14.	<u>1</u> CH <sub>4</sub> + <u>2</u> O <sub>2</sub> → <u>1</u> CO <sub>2</sub> + 2 H <sub>2</sub> O		

	<b>#14</b>	<b>Before</b>	<b>After</b>
15. a)	<u>  </u> ZnCl <sub>2</sub> + <u>  </u> (NH <sub>4</sub> ) <sub>2</sub> S → <u>  </u> ZnS + <u>  </u> NH <sub>4</sub> Cl		
b)	Find the molar mass of these reactants.		
c)	How many moles of ZnCl <sub>2</sub> would be in 25 g? How much mass would 0.55 moles of (NH <sub>4</sub> ) <sub>2</sub> S have?		

# Chemistry Unit 7 Test Chemical Reaction Answers

**Shasha Hu**

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