

Moles and Reactions

Name: _____

Date: _____

Hour: _____

Information: Mole Ratios in Equations

Propane is burned in many rural homes for heat in the winter. Below is the balanced equation for the combustion of propane (C_3H_8).



For each molecule of propane that is burned, there needs to be five molecules of oxygen present. Likewise, if there were a dozen molecules of propane, five dozen molecules of oxygen would be required. Similarly, for each **mole** of propane, five **moles** of oxygen are needed. Also, for each mole of propane burned three moles of carbon dioxide and 4 moles of water are produced. **The numbers of moles of each substance in a chemical equation are related by the ratio of the coefficients of each substance.**

Critical Thinking Questions

Note: For questions 1-6, refer to the balanced equation for the combustion of propane.

1. a) How many moles of water are produced when 1.45 moles of propane are combusted?

It is a 4:1 ratio so the moles of water = $(1.45)(4) = 5.8$ moles

- b) How many molecules of water is this? (Remember each mole has 6.02×10^{23} molecules.)

$(5.8)(6.02 \times 10^{23}) = 3.49 \times 10^{24}$

2. If 2.35 moles of CO_2 are produced in a reaction, how many moles of H_2O would be produced?

It is a 4:3 ratio so the moles of water = $2.35(4/3) = 3.13$ moles

3. Why is this statement false: "If 10 grams of propane burn, you need 50 grams of oxygen."

The coefficients in the balanced equation relate mole ratios and not gram ratios. The above statement would be true if it read: "If 10 moles of propane burn, you need 50 moles of oxygen."

4. a) If 27.3 moles of carbon dioxide are produced during the combustion of a certain amount of propane, how many moles of propane were combusted?

Since it is a 1:3 ratio, the moles of propane = $27.3(1/3) = 9.1$ moles

- b) How many grams of propane was this?

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Table of Contents Chemquest 32 Moles And Reactions Answer Key

1. Understanding the eBook Chemquest 32 Moles And Reactions Answer Key
 - The Rise of Digital Reading Chemquest 32 Moles And Reactions Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemquest 32 Moles And Reactions Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemquest 32 Moles And Reactions Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemquest 32 Moles And Reactions Answer Key
 - Personalized Recommendations
 - Chemquest 32 Moles And Reactions Answer Key User Reviews and Ratings
 - Chemquest 32 Moles And Reactions Answer Key and Bestseller Lists

5. Accessing Chemquest 32 Moles And Reactions Answer Key Free and Paid eBooks
 - Chemquest 32 Moles And Reactions Answer Key Public Domain eBooks
 - Chemquest 32 Moles And Reactions Answer Key eBook Subscription Services
 - Chemquest 32 Moles And Reactions Answer Key Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Chemquest 32 Moles And Reactions Answer Key Enhanced eBook Features
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 - Highlighting and Note-Taking Chemquest 32 Moles And Reactions Answer Key
 - Interactive Elements Chemquest 32 Moles And Reactions Answer Key
8. Staying Engaged with Chemquest 32 Moles And Reactions Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemquest 32 Moles And Reactions Answer Key
9. Balancing eBooks and Physical Books Chemquest 32 Moles And Reactions Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemquest 32 Moles And Reactions Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemquest 32 Moles And Reactions Answer Key
 - Setting Reading Goals Chemquest 32 Moles And Reactions Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemquest 32 Moles And Reactions Answer Key
 - Fact-Checking eBook Content of Chemquest 32 Moles And Reactions Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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