



balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay

electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- Sulfur hexafluoride SF_6
- Lithium nitride Li_3N
- Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- Tin (II) chloride SnCl_2
- Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- Mercury (I) chloride Hg_2Cl_2
- Potassium bromate KBrO_3
- Hydrosulfuric acid $\text{H}_2\text{S}(\text{aq})$
- Chloric acid HClO_3
- Sulfurous acid H_2SO_3

Name each of the following compounds.

- CuSO_4 Copper (II) sulfate
- AlF_3 aluminum fluoride
- HI hydrogen iodide
- NO nitrogen monoxide
- H_2Se hydroseonic acid
- HNO_3 nitric acid
- HNO_2 nitrous acid
- NaHSO_4 sodium bisulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 43 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

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Arshad Iqbal

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Table of Contents Chemistry Study Guide With Answers

1. Understanding the eBook Chemistry Study Guide With Answers
 - The Rise of Digital Reading Chemistry Study Guide With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide With Answers
 - Personalized Recommendations

- Chemistry Study Guide With Answers User Reviews and Ratings
- Chemistry Study Guide With Answers and Bestseller Lists

5. Accessing Chemistry Study Guide With Answers Free and Paid eBooks

- Chemistry Study Guide With Answers Public Domain eBooks
- Chemistry Study Guide With Answers eBook Subscription Services
- Chemistry Study Guide With Answers Budget-Friendly Options

6. Navigating Chemistry Study Guide With Answers eBook Formats

- ePUB, PDF, MOBI, and More
- Chemistry Study Guide With Answers Compatibility with Devices
- Chemistry Study Guide With Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Chemistry Study Guide With Answers
- Highlighting and Note-Taking Chemistry Study Guide With Answers
- Interactive Elements Chemistry Study Guide With Answers

8. Staying Engaged with Chemistry Study Guide With Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Study Guide With Answers

9. Balancing eBooks and Physical Books Chemistry Study Guide With Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Chemistry Study Guide With Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Chemistry Study Guide With Answers

- Setting Reading Goals Chemistry Study Guide With Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Study Guide With Answers

- Fact-Checking eBook Content of Chemistry Study Guide With Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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