

Derive the following equations:

1. Show the following thermodynamic equations:

$$d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -d(\Delta nRT)$$

(a) is this equation a thermodynamic or mathematical relationship only?

$$d(\Delta nRT) = d(\Delta H_{\text{mix}}) - Td(\Delta S_{\text{mix}}) + d(\Delta PV_{\text{mix}})$$

(b) what equation do we need if both ΔH_{mix} and ΔPV_{mix} were known?

$$d(\Delta nRT) = \frac{d(\Delta H_{\text{mix}})}{T} + \frac{d(\Delta PV_{\text{mix}})}{T} - d(\Delta S_{\text{mix}})$$

(c) what equation do we need if both ΔS_{mix} and ΔPV_{mix} were known?

$$\Delta nRT = \frac{\Delta H_{\text{mix}}}{T} + \frac{d(\Delta PV_{\text{mix}})}{T} - \frac{d(\Delta S_{\text{mix}})}{T}$$

(d) what does ΔPV_{mix} represent for a mixture of gases (ideal gases)?

$$\Delta nRT = \frac{\Delta H_{\text{mix}}}{T} + \frac{d(\Delta PV_{\text{mix}})}{T} - \frac{d(\Delta S_{\text{mix}})}{T}$$

2. Thermodynamic equations: $d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -d(\Delta nRT)$

all of the thermodynamic equations which must include the appropriate d with the appropriate sign for the differentiation of thermodynamic equation

$$d(\Delta nRT) = d(\Delta H_{\text{mix}}) - Td(\Delta S_{\text{mix}}) + d(\Delta PV_{\text{mix}}) \quad d(\Delta nRT) = d(\Delta H_{\text{mix}}) - Td(\Delta S_{\text{mix}}) + d(\Delta PV_{\text{mix}})$$

3. If a single substance (pure) was compressed, conditions would be identical or related to the standard state conditions: $d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -d(\Delta nRT)$

$$d(\Delta nRT) = \frac{d(\Delta H_{\text{mix}})}{T} + \frac{d(\Delta PV_{\text{mix}})}{T} - d(\Delta S_{\text{mix}})$$

4. If a single substance (pure) was compressed by the equation of state (ideal gas):

$$d(\Delta nRT) = \frac{d(\Delta H_{\text{mix}})}{T} + \frac{d(\Delta PV_{\text{mix}})}{T} - d(\Delta S_{\text{mix}})$$

5. If a single substance (pure) was compressed by the equation of state (ideal gas) and the standard state was known:

$$d(\Delta nRT) = \frac{d(\Delta H_{\text{mix}})}{T} + \frac{d(\Delta PV_{\text{mix}})}{T} - d(\Delta S_{\text{mix}})$$

6. If a single substance (pure) was compressed, $d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -d(\Delta nRT)$

a. Show a balanced equation for the reaction:



b. If a single substance (pure) was compressed, $d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -d(\Delta nRT)$

$$\frac{d(\Delta H_{\text{mix}})}{T} + \frac{d(\Delta PV_{\text{mix}})}{T} - d(\Delta S_{\text{mix}})$$

c. If a single substance (pure) was compressed, $d(\Delta G_{\text{mix}}) = Td(\Delta S_{\text{mix}}) - d(\Delta H_{\text{mix}}) + d(\Delta PV_{\text{mix}}) = -d(\Delta nRT)$

$$d(\Delta nRT) = d(\Delta H_{\text{mix}}) - Td(\Delta S_{\text{mix}}) + d(\Delta PV_{\text{mix}})$$

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George Facer



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