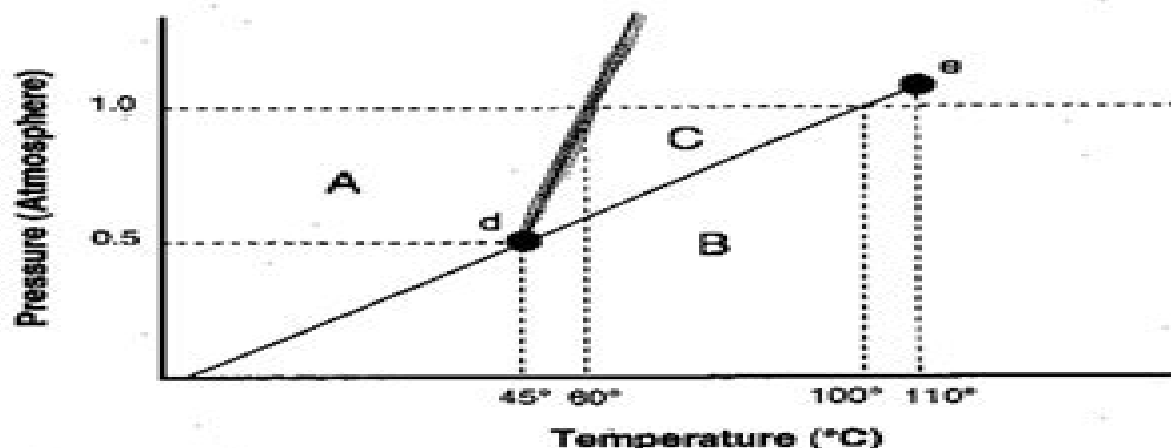


PHASE DIAGRAM

Name Key



Answer the following questions using the chart above.

1. What section represents the solid phase? A
2. What section represents the liquid phase? C
3. What section represents the gas phase? B
4. What letter represents the triple point? d
5. What letter represents the critical point? e
6. What is this substance's ^{1 atm}normal melting point? 60°C
7. What is this substance's ^{1 atm}normal boiling point? 100°C
8. Above what temperature is it impossible to liquify this substance no matter what the pressure? 110°C
9. At what temperature and pressure do all three phases coexist? 45°C
10. Is the density of the solid greater than or less than the density of the liquid?
greater b/c (+) slope
11. Would an increase in pressure cause this substance to freeze or melt? freeze

Chemistry Ws Phase Diagrams Answers

**Thomas Williams Bicknell, Albert
Edward Winship, Anson Wood Belding**



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