

## Unit 5 Study Guide

2. In soda pop,  $\text{CO}_2(\text{g})$  is a \_\_\_\_\_, and water is the \_\_\_\_\_.

- A) solute; solvent
- B) solvent; solute
- C) solution; solute
- D) solute; solution
- E) solvent; solution

Ans: A

3. A mixture of sand and water is a(n) \_\_\_\_\_.

- A) solution
- B) solvent
- C) solute
- D) aqueous solution
- E) none of these

Ans: E

4. True or false? Approximately 38 g of NaCl can be dissolved in 100 g of water at 25°C. A solution prepared by adding 35 g of NaCl to 100 g of water at 25°C is unsaturated.

- A) True
- B) False

Ans: A

5. When a solvent has dissolved all the solute it can at a particular temperature, it is said to be

- A) dilute
- B) unsaturated
- C) supersaturated
- D) saturated
- E) none of these

Ans: D

8. An oven-cleaning solution is 40.0% (by mass) NaOH. If one jar of this product contains 457 g of solution, how much NaOH does it contain?

- A)  $1.14 \times 10^3$  g
- B) 11.4 g
- C) 183 g
- D) 18.3 g
- E) none of these

Ans: C

9. A nitric acid solution containing 71.0%  $\text{HNO}_3$  (by mass) has a density of 1.42 g/ml. How many moles of  $\text{HNO}_3$  are present in 2.02 L of this solution?

- A) 64.1 mol
- B) 32.3 mol
- C) 22.8 mol
- D)  $3.09 \times 10^{-2}$  mol
- E) none of these

Ans: B

13. Determine the concentration of a solution made by dissolving 24.8 g of sodium chloride in 750.0 mL of solution.

- A) 0.318 M
- B) 33.1 M
- C) 0.424 M
- D) 0.566 M
- E) 18.6 M

Ans: D

14. If you mix 20.0 mL of a 3.00 M sugar solution with 30.0 mL of a 4.11 M sugar solution, you will end up with a sugar solution of \_\_\_\_\_.

- A) 3.55 M
- B) 0.142 M
- C) 6.11 M
- D) 3.67 M
- E) 9.17 M

Ans: D

15. You have 250.0 mL of 4.00 M sugar solution. You add 217.3 mL of water to this solution. Determine the concentration of the sugar solution after the water has been added.

- A) 4.00 M
- B) 2.39 M
- C) 2.88 M
- D) 2.14 M
- E) 2.47 M

Ans: D

17. Which of the following aqueous solutions contains the greatest number of ions in solution?

- A) 2.0 L of 1.50 M sodium phosphate
- B) 2.0 L of 1.50 M magnesium chloride
- C) 3.0 L of 1.50 M sodium chloride
- D) 2.0 L of 2.00 M potassium fluoride
- E) 1.0 L of 3.00 M sodium sulfate

Ans: A

19. What is the minimum volume of a 2.42 M NaOH solution needed to make 150.0 mL of a 0.3000 M NaOH solution?

- A) 375 mL
- B) 24.8 mL
- C) 61.6 mL
- D) 120. mL
- E) 49.6 mL

Ans: E

# Chemistry Study Guide Solution Concentration Answers

**Martin Silberberg, Libby Weberg**

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