

ChemQuest 35

Intro. to Gases

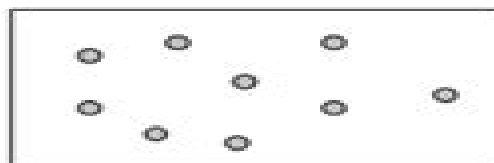
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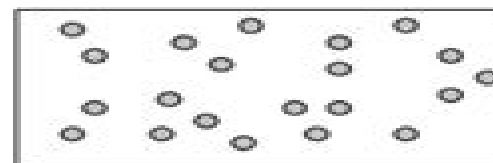
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Information: Gas Pressure

Figure 1: Two containers of gas molecules



Container 1



Container 2

Gas molecules move randomly in their containers, colliding with the walls of their containers causing "gas pressure." Pressure can be defined as the force pushing on an area. It can be described with the equation:

$$P = \frac{F}{A} \quad \text{where } P \text{ is pressure (in kPa), } F \text{ is force (in N) and } A \text{ is area (in m}^2\text{).}$$

The more molecules collide with the wall and the faster the molecules are going when they strike the wall, the greater the force on the wall and therefore, the higher the pressure.

Critical Thinking Questions

1. Use the pressure equation to explain why it would be more likely for an ice skater to fall through the ice on a lake than it would be for someone walking across the lake with regular shoes on.
The blade of an ice skate has a much smaller surface area than the sole of a shoe. With a smaller area (A), the pressure (P) on the ice is larger.
2. Which container in Figure 1 has the highest pressure? Explain.
Container 2, because there are more gas molecules just like a tire filled with more air.
3. If I heated container 1 and did not heat container 2, could I get the pressure in container 1 to equal container 2? Explain.
Yes, increasing the temperature increases pressure by speeding up gas molecules.
4. If container 2 was made of an elastic material and if I expanded container 2, could I make the two containers have equal gas pressures? Explain.
Yes, increasing volume decreases pressure giving both containers an equal number of molecules per volume.

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AW Chickering



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