

## Chapter Goals 1.1

# Intro. to Chemical Reactions

Name: \_\_\_\_\_  
 Date: \_\_\_\_\_  
 Class: \_\_\_\_\_

## Electrolyte Review

1. Write the formula for the following compounds:

- a) barium phosphate:  $\text{Ba}_3(\text{PO}_4)_2$       b) sodium phosphate:  $\text{Na}_3\text{PO}_4$   
 c) magnesium phosphate:  $\text{Mg}_3(\text{PO}_4)_2$       d) lithium nitrate:  $\text{LiNO}_3$

2. Write the name for the following compounds:

- a)  $\text{Cu}(\text{NO}_3)_2$ : copper(II) nitrate      b)  $\text{FeCl}_3$ : iron(III) chloride  
 c)  $\text{SnSO}_4$ : tin(II) sulfate      d)  $\text{SnF}_4$ : tin(IV) fluoride

## Chemical Reactions: What they look like...

Here is what a chemical reaction looks like when it is written:



There are examples that you can write to describe what happens above:

Chemical reactions result with atoms and molecules and ions and net ionic

There are also some characteristics for chemical reactions and the reactions themselves. Here are a few points:

## Critical Thinking Questions

1. Write the following reaction, with the reactants and products:



Chemical reactions result with atoms and molecules and ions and net ionic

# Chemquest 22 Answer Key

**A Gutmann**



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