



CHEMISTRY

CHEAT SHEET

GAS LAWS



GAS LAWS CHEAT SHEET

Remember: a gas is a substance composed of particles which move from one volume to another (diffusion) and which are separated by empty space. Remember: a substance is a material with a definite composition and a definite chemical formula. Remember: a gas is a substance which is not condensed into a liquid or solid at room temperature and pressure.

Boyle's Law
Charles' Law
Avogadro's Law

Boyle's Law	Charles' Law	Avogadro's Law
Pressure (P) is inversely proportional to Volume (V) at constant Temperature (T).	Volume (V) is directly proportional to Temperature (T) at constant Pressure (P).	Volume (V) is directly proportional to the number of moles (n) of gas at constant Temperature (T) and Pressure (P).
$P \propto \frac{1}{V}$	$V \propto T$	$V \propto n$
$P_1 V_1 = P_2 V_2$	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$

Boyle's Law

Remember: pressure and volume are inversely proportional.

$$P_1 V_1 = P_2 V_2$$

Remember: if you increase the pressure of a gas, the volume will decrease, and vice versa.

Charles' Law

Remember: volume and temperature are directly proportional.

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Remember: if you increase the temperature of a gas, the volume will increase, and vice versa.

Avogadro's Law

Remember: volume and the number of moles are directly proportional.

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

Remember: if you increase the number of moles of a gas, the volume will increase, and vice versa.

Ideal Gas Law

The ideal gas law is a combination of Boyle's, Charles', and Avogadro's laws. It is written as follows:

$$PV = nRT$$

Where:

- P = pressure (atm)
- V = volume (L)
- n = number of moles (mol)
- R = gas constant (0.0821 L·atm/mol·K)
- T = temperature (K)

Remember: if you increase the pressure of a gas, the volume will decrease, and vice versa.

$P \propto \frac{nT}{V}$	$V \propto \frac{nT}{P}$
$n \propto \frac{PV}{RT}$	$T \propto \frac{PV}{nR}$

Chemistry Study Guide Gas Laws

Arshad Iqbal



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