

Unit 1 Chemistry Study Guide

History of the Atom:

Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
- **Dalton's Atomic Theory:**
 - o All matter is composed of extremely small particles called atoms.
 - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
 - o Atoms cannot be subdivided, created, or destroyed.
 - o Atoms of different elements can combine in simple whole number ratios to form chemical compounds.
 - o In chemical reactions, atoms are combined, separated or rearranged.
- **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model:



J.J. Thompson (1906):

- **Cathode Ray Experiment –**
 - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

Chemistry Study Guide Nuclear Chemistry

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Radiochemistry and Nuclear Chemistry Gregory Choppin, Jan-Olov Liljenzin, Jan Rydberg, 2002 Origin of Nuclear Science Nuclei Isotopes and Isotope Separation Nuclear Mass and Stability Unstable Nuclei and Radioactive Decay Radionuclides in Nature Absorption of Nuclear Radiation Radiation Effects on Matter Detection and Measurement Techniques Uses of Radioactive Tracers Cosmic Radiation and Elementary Particles Nuclear Structure Energetics of Nuclear Reactions Particle Accelerators Mechanics and Models of Nuclear Reactions Production of Radionuclides The Transuranium Elements Thermonuclear Reactions the Beginning and the Future Radiation Biology and Radiation Protection Principles of Nuclear Power Nuclear Power Reactors Nuclear Fuel Cycle Behavior of Radionuclides in the Environment Appendices Solvent Extraction Separations Answers to Exercises Isotope Chart Periodic Table of the Elements Quantities and Units Fundamental Constants Energy Conversion Factors Element and Nuclide Index Subject Index

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Radiochemistry and Nuclear Chemistry Gregory Choppin,Jan-Olov Liljenzin,Jan Rydberg,Christian Ekberg,2013-09-05

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