

**CHAPTER 4 PRACTICE TEST: SOLUTION STOICHIOMETRY**

1) Which of the following are weak electrolytes? HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$ , KCl  
 A) HCl, KCl  
 B) HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$ , KCl  
 C)  $\text{HC}_2\text{H}_3\text{O}_2$ , KCl  
 D)  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$   
 E) HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ , KCl

Answer: D Diff: 2

2) What are the spectator ions in the reaction between KOH (aq) and  $\text{HNO}_3$  (aq)?

A)  $\text{K}^+$  and  $\text{H}^+$       C)  $\text{K}^+$  and  $\text{NO}_3^-$       E)  $\text{OH}^-$  only  
 B)  $\text{H}^+$  and  $\text{OH}^-$       D)  $\text{H}^+$  and  $\text{NO}_3^-$

Answer: C Diff: 2

3) The balanced net ionic equation for precipitation of  $\text{CaCO}_3$  when aqueous solutions of  $\text{Na}_2\text{CO}_3$  and  $\text{CaCl}_2$  are mixed is \_\_\_\_\_.

A)  $2\text{Na}^+$  (aq) +  $\text{CO}_3^{2-}$  (aq)  $\rightarrow$   $\text{Na}_2\text{CO}_3$  (aq)  
 B)  $2\text{Na}^+$  (aq) +  $2\text{Cl}^-$  (aq)  $\rightarrow$   $2\text{NaCl}$  (aq)  
 C)  $\text{Na}^+$  (aq) +  $\text{Cl}^-$  (aq)  $\rightarrow$   $\text{NaCl}$  (aq)  
 D)  $\text{Ca}^{2+}$  (aq) +  $\text{CO}_3^{2-}$  (aq)  $\rightarrow$   $\text{CaCO}_3$  (s)

Answer: D Diff: 3

4) Which of the following are strong acids? HI,  $\text{HNO}_3$ , HF, HBr

A) HF, HBr      C) HI, HF, HBr      E) HI,  $\text{HNO}_3$ , HBr  
 B) HI,  $\text{HNO}_3$ , HF, HBr      D)  $\text{HNO}_3$ , HF, HBr

Answer: E Diff: 2

5) A neutralization reaction between an acid and a metal hydroxide produces \_\_\_\_\_.

A) water and a salt      C) oxygen gas      E) ammonia  
 B) hydrogen gas      D) sodium hydroxide

Answer: A Diff: 2

6) Which of these metals is the least easily oxidized? Na, Au, Fe, Ca, Ag

A) Na      B) Au      C) Fe      D) Ca      E) Ag

Answer: B Diff: 3

7) Based on the equations below, which metal is the most active?



A) Ni      B) Ag      C) Cu      D) Pb      E) N

Answer: A Diff: 3

8) What is the concentration (M) of KCl in a solution made by mixing 25.0 mL of 0.100 M KCl with 50.0 mL of 0.100 M KCl?

A) 0.100      B) 0.0500      C) 0.0333      D) 0.0250      E) 125

Answer: A Diff: 3

# Chemistry Test Answer Key

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(Jr), Henry Hurd Rusby, Daniel Base**

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