

COMMON CHEMISTRY TERMS

Matter

Atom - basic unit of an element
Atom - negatively charged ion
Compound - group of two or more different atoms bonded together (H_2O , CO_2 , CH_4)
Cation - positively charged ion
Chemistry - the study of matter
Diatomic - two highly bonded atoms in pairs (= diatomic)
Electron - tiny negative electrically charged particles of an atom
Ions - charged particles due to loss or gain of electrons
Isomers - multiple forms of same chemical with different molecular structures
Matter - anything that takes up space and has mass
Molecule - group of atoms bonded together (H_2 , O_2 , H_2O , CO_2)
Nucleus - central of an atom consisting of protons and neutrons
Neutron - neutral particle in the nucleus of an atom
Phase - a form of matter solid, liquid, gas, or plasma
Protons - positively charged inside nucleus

Chemical Reaction Vocabulary

Catalyst - something that speeds up the rate of a chemical reaction
Endothermic - a reaction that absorbs heat
Entropy - degree of disorder in a system
Enzyme - a biological catalyst
Equilibrium - when the rate of the forward reaction equals the rate of the reverse reaction
Exothermic - a reaction that releases heat
Reactants - substances that undergo a change in a reaction (on the left of the arrow in a chemical equation)
Preproduct - a solid that forms from a solution
Products - substance that is formed in a chemical reaction (on the right of the arrow in a chemical equation)

Common Types of Reactions

Combustion: $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ - exothermic
Decomposition: $\text{AB} \rightarrow \text{A} + \text{B}$
Dehydration - a reaction that results in loss of water (water appears in the products)
Double Displacement: $\text{AB} + \text{CD} \rightarrow \text{AD} + \text{CB}$
Hydration - a reaction where a chemical substance reacts with water to form a product (water is in the reactants)
Neutralization: $\text{Acid} + \text{Base} \rightarrow \text{H}_2\text{O} + \text{salt}$
Synthesis: $\text{A} + \text{B} \rightarrow \text{AB}$
Single Displacement: $\text{A} + \text{BC} \rightarrow \text{AC} + \text{B}$

Solutions, Acids, and Bases

Acid - molecule that can either donate a proton (H^+) or form a constant bond by accepting a pair of electrons. Types include:
dicarboxylic acids - donate a proton
Arrhenius acid - forms constant bond by accepting electron pair
Amphiprotic - substance that can react as an acid and a base
Base - molecule that can increase concentration of hydroxide (OH^-) or accept a proton (H^+) or donate a pair of electrons. Types include:
Arrhenius Base - increases concentration of hydroxide ions
dicarboxylic bases - accept a proton
Arrhenius base - donate an electron pair
Concentration - how much of a substance is mixed with another substance
Molarity - the number of moles of a substance in a certain volume of solution. An example of a concentration
Mole - standard unit used to measure the amount of any substance (1 mole = 6.022 x 10²³ individual components)
pH - a measure of the acidity of a solution
Solute - what is dissolved in the solvent
Solvent - homogeneous mixture with 7 or more components
Solvent - what a solute is dissolved in

Periodic Trends

Atomic radius - size of an atom
Electron Affinity - amount of energy released when an atom accepts an electron
Electronegativity - ability of an atom to attract shared electrons in bond
Ionization Energy - amount of energy required to remove an electron from an atom

Bonding + Intermolecular Forces

Covalent Bond - a bond where electrons are shared between atoms
Dipole-dipole forces - attractive intermolecular forces between positive end of a polar molecule and the negative end of another polar molecule
Hydrogen bonding - a type of dipole-dipole intermolecular force that occurs with the following bond types: H-F , H-O , and H-N
Intermolecular force - the forces that occur between separate atoms and molecules
Ionic Bond - bond formed by the transfer of electrons between atoms
Lewis Structure - simplified representation of valence electrons for an atom
Lewis dot structure - valence intermolecular force between atoms and molecules
Valence electrons - electrons in outermost shell of an atom

Chemistry Vocabulary Study Guide

Lewis Morris



Chemistry Vocabulary Study Guide:

Chemistry Regents Vocabulary Workbook Lewis Morris, Learn the Secret to Success on the Chemistry Regents Exam Ever wonder why learning comes so easily to some people This remarkable workbook reveals a system that shows you how to learn faster easier and without frustration By mastering the hidden language of the subject and exams you will be poised to tackle the toughest of questions with ease We ve discovered that the key to success on the Chemistry Regents Exam lies with mastering the Insider s Language of the subject People who score high on their exams have a strong working vocabulary in the subject tested They know how to decode the vocabulary of the subject and use this as a model for test success People with a strong Insider s Language consistently Perform better on their Exams Learn faster and retain more information Feel more confident in their courses Perform better in upper level courses Gain more satisfaction in learning The Chemistry Regents Exam Vocabulary Workbook is different from traditional review books because it focuses on the exam s Insider s Language It is an outstanding supplement to a traditional review program It helps your preparation for the exam become easier and more efficient The strategies puzzles and questions give you enough exposure to the Insider Language to use it with confidence and make it part of your long term memory The Chemistry Regents Exam Vocabulary Workbook is an awesome tool to use before a course of study as it will help you develop a strong working Insider s Language before you even begin your review Learn the Secret to Success After nearly 20 years of teaching Lewis Morris discovered a startling fact Most students didn t struggle with the subject they struggled with the language It was never about brains or ability His students simply didn t have the knowledge of the specific language needed to succeed Through experimentation and research he discovered that for any subject there was a list of essential words that when mastered unlocked a student s ability to progress in the subject Lewis called this set of vocabulary the Insider s Words When he applied these Insider s Words the results were incredible His students began to learn with ease He was on his way to developing the landmark series of workbooks and applications to teach this Insider s Language to students around the world

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