

balancing Equations

Name: _____

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Introduction: Balancing

- Write the equation for the reaction of sodium and chlorine (chlorine gas) to form sodium chloride.



- Write the equation for the reaction of sodium oxide and carbon dioxide to produce sodium carbonate sodium.



Information: Subscripts and Coefficients

It is important to understand that atoms are never really "destroyed or produced". For example, in Cl_2 there are two chlorine atoms and when this is broken down into $2Cl$ there are still two chlorine atoms.

A coefficient tells you how many molecules (or atoms) there are, that is a different way. For example in the equation: " $2H_2O$ " the 2 here is the coefficient. The 2 here tells us that there are three molecules of water present. In this equation " $2H_2O$ " there are a total of 4 hydrogen atoms and 2 oxygen atoms.

Balancing Chemical Equations

- How many oxygen atoms are present in $2H_2O$? $2 \times 1 = 2$
- How many oxygen atoms are in each of the following:



Information: How To Balance Equations

Consider the reaction of sodium and chlorine to produce sodium chloride (halogenation):

$$Na + Cl_2 \rightarrow NaCl$$

Remember that when chlorine is in Cl_2 it is always written as Cl_2 . On the product side of the equation there only there are a total of two atoms of chlorine, but on the product side there is only one atom of chlorine. Atomic masses always disagree. To make the equation for make sense, we need to balance the equation. This can be done. But, by adding a "2" to the product side:



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The Directory of Management Consultants, 2003 Kennedy Information (Firm), 2002

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