



#### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay

#### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



#### NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

a. Sulfur hexafluoride  $\text{SF}_6$

b. Lithium nitride  $\text{Li}_3\text{N}$

c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$

d. Tin (II) chloride  $\text{SnCl}_2$

e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )

f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$

g. Potassium bromate  $\text{KBrO}_3$

h. Hydrosulfuric acid  $\text{H}_2\text{S(OH)}_2$

i. Chloric acid  $\text{HClO}_3$

j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds.

a.  $\text{CuSO}_4$  Copper (II) sulfate

b.  $\text{AlF}_3$  aluminum fluoride

c.  $\text{HI}$  hydrogen iodide

d.  $\text{NO}$  nitrogen monoxide

e.  $\text{H}_2\text{Se}$  hydrogen selenide

f.  $\text{HNO}_3$  nitric acid

g.  $\text{HNO}_2$  nitrous acid

h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

#### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it takes for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

#### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 43 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

# Chemistry Study Guide With Answers

**Didier Musso**

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*Organic Chemistry, Study Guide and Solutions Manual* T. W. Graham Solomons,Craig B. Fryhle,2000

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**Chemistry** Steven S. Zumdahl,2000

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[Student Study Guide for Chemistry](#) John E. McMurry, Robert C. Fay, Jill Kirsten Robinson, Julie Klare, Donnajean Fredeen, James Zubricky, 2015-03-09

The Study Guide includes learning goals, an overview, a review section with worked examples and self tests with answers.

**Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan

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*Introduction to Chemistry* Mark Bishop, 2001-12

**ACS General Chemistry Study Guide** Joshua

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## **Table of Contents Chemistry Study Guide With Answers**

1. Understanding the eBook Chemistry Study Guide With Answers
  - The Rise of Digital Reading Chemistry Study Guide With Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide With Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Study Guide With Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide With Answers
  - Personalized Recommendations
  - Chemistry Study Guide With Answers User Reviews and Ratings
  - Chemistry Study Guide With Answers and Bestseller Lists
5. Accessing Chemistry Study Guide With Answers Free and Paid eBooks

- Chemistry Study Guide With Answers Public Domain eBooks
- Chemistry Study Guide With Answers eBook Subscription Services
- Chemistry Study Guide With Answers Budget-Friendly Options

6. Navigating Chemistry Study Guide With Answers eBook Formats

- ePUB, PDF, MOBI, and More
- Chemistry Study Guide With Answers Compatibility with Devices
- Chemistry Study Guide With Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Chemistry Study Guide With Answers
- Highlighting and Note-Taking Chemistry Study Guide With Answers
- Interactive Elements Chemistry Study Guide With Answers

8. Staying Engaged with Chemistry Study Guide With Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Study Guide With Answers

9. Balancing eBooks and Physical Books Chemistry Study Guide With Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Chemistry Study Guide With Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Chemistry Study Guide With Answers

- Setting Reading Goals Chemistry Study Guide With Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Study Guide With Answers

- Fact-Checking eBook Content of Chemistry Study Guide With Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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