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web bonus problem here is a balanced net ionic equation $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}^- + 3\text{H}_2\text{O}$ the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + 6\text{Fe}^{2+} + 14\text{H}^+ \rightarrow 2\text{Cr}^{3+} + 6\text{Fe}^{3+} + 7\text{H}_2\text{O}$ solution 1 balanced half reactions 6e 14h $\text{Cr}_2\text{O}_7^{2-} + 6\text{Fe}^{2+} + 14\text{H}^+ \rightarrow 2\text{Cr}^{3+} + 6\text{Fe}^{3+} + 7\text{H}_2\text{O}$ equalize the electrons 6e 14h $\text{Cr}_2\text{O}_7^{2-} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O} \rightarrow 6\text{Fe}^{3+} + 6\text{Fe}^{2+} + 6\text{H}^+$ multiplied by a factor of 6

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web balance each half reaction in basic solution 4 5 6 7 $\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + 6\text{Cl}^- \rightarrow 2\text{Cr}^{3+} + 3\text{Cl}_2 + 7\text{H}_2\text{O}$ no $\text{SO}_4^{2-} + 2\text{MnO}_4^- + 8\text{H}^+ \rightarrow 2\text{Mn}^{2+} + 2\text{H}_2\text{O} + \text{O}_2 + 4\text{H}_2\text{O}$ balance each redox reaction in acid solution using the half reaction method 8 $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 6\text{H}^+ \rightarrow 2\text{Cr}^{3+} + 3\text{O}_2 + 8\text{H}_2\text{O}$ 9 $\text{TeO}_3 + 2\text{H}_2\text{O} \rightarrow \text{TeO}_4^- + 4\text{H}^+$ 10 $\text{MnO}_4^- + 3\text{H}_2\text{O} \rightarrow \text{MnO}_2 + 5\text{H}_2\text{O}$ 11 $\text{PbO}_2 + 4\text{H}^+ + 2\text{Cl}^- \rightarrow \text{Pb}^{2+} + \text{Cl}_2 + 2\text{H}_2\text{O}$ 12 $\text{MnO}_4^- + 8\text{H}^+ + 5\text{Cl}^- \rightarrow \text{Mn}^{2+} + 5\text{Cl}_2 + 4\text{H}_2\text{O}$ as $\text{NO}_3^- + 4\text{H}^+ + 3\text{I}^- \rightarrow \text{NO} + \text{I}_2 + 2\text{H}_2\text{O}$ re $\text{IO}_3^- + 6\text{H}^+ + 5\text{Cl}^- \rightarrow \text{I}_2 + \text{Cl}_2 + 3\text{H}_2\text{O}$ h₂aso₄

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web practice exercises balance the following equations of redox reactions assign oxidation numbers to all elements in the reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding h 2 o h and e

in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h 2no 2 c 6 h 6 o 6 2h 2 o 2no choose 1 answer c 6 h 8 o 6 and no 2 a c 6 h 8 o 6 and no 2 no 2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h and