

Acids and Bases

Overview of Acids and Bases

Indicators: Acids and Bases

Acids: Donates its proton to base (proton donor)

Base: Donates its lone pair to form a bond (LBP)



Remember: (1) only H atoms, (2) only OH is a base

What about organic indicators and how organic indicators?

Acids and Bases

Acids are proton donors, (the electrons are being to move)
Bases are proton acceptors, (the electrons are being to move)



General acid-base reactions of acid-base reactions:

to find the strong base, the base gets negative charge of the base

and check for weak acids hydrogens of the acid.

to find the weak acid, the base gets negative charge of the base

in the other case.

Remember: -OH is a weak base, -OH is a weak acid.

1. Acid: Donates its proton to base (proton donor)

2. Base: Accepts its proton to base (proton acceptor)

Acids and Bases

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Acid: Donates its proton to base (proton donor)

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Acid Strength

What does it mean "strong acid"?

Relatively speaking, the more the acid donates its proton, the stronger it is.



The dissociation of the acid is a reversible reaction that is described by the equilibrium constant.



Acid Strength and pKa

The Ka value of an acid (acid) and its conjugate base (B) are shown below.



The acid is a stronger acid than its conjugate base.

To compare the acids, the pKa is used, which is the negative log of the Ka.

$$\text{pKa} = -\log K_a$$

$$\text{pKa}_1 = -\log K_{a1} = 10$$

$$\text{pKa}_2 = -\log K_{a2} = 10$$

So, the stronger the acid, the lower the pKa value. (It's the opposite of the Ka)

Acids are pKa values of their conjugate base (pKa)



What makes the Acid Strong?

The strength of the acid is determined by the stability of the conjugate base (the base).

The stronger the acid, the more stable the conjugate base.



Acids are pKa values of their conjugate base (pKa)

Base is the base of the conjugate acid (pKa)

1. The first and last factor is the **charge** of the conjugate base (the base).

to make the conjugate base (the base) more stable.



2. When comparing charges in different ways, the more stable the base.

Large atoms have higher polarizability and results for negative charge better in the charge is spread out.



There are weak acids and strong acids in more electronegative than water.

The water atom is large and stabilizes the negative charge better.

3. When comparing for more atoms, check which conjugate base is **more electronegative**.



1. The first factor for determining the acidity is called **inductive**.

Inductive is inductive effect is the transmission of charge through atoms.

Electronegativity atoms pull the negative charge and helps to stabilize the conjugate base.



The more electronegative, the more electron withdrawing, the stronger the acid.



The closer the electronegative atom to the hydrogen, the stronger the acid.

Electronegativity

2. The second factor is the **charge** of the conjugate base (the base).

The negative charge difference of pKa values of acids, bases, and bases.



3. Inductive is more electronegative than H, so the more H character, the stronger the acid.

Therefore, for **acids**, **inductive** = electron withdrawing.

The primary factor, **charge**, **inductive**, **charge**, **charge**, **charge**, **charge**.

to a conjugate base (the base).

Remember: the stronger the conjugate base.



The Position of Equilibrium for an Acid-Base Reaction

to the following reaction, which is possible?



Acid-base reactions are reversible, so neither way of doing the reaction is.

to the equilibrium of the reaction, which is the right (products) or the left (reactants)?

To answer this question, you need to remember: for the equilibrium of acids and bases.

reaction is shifted towards the weaker acid and the base.

So, you need to compare the pKa of the acid and the pKa of the conjugate acid.

The equilibrium will be shifted to the higher pKa reaction with only.



to the acid, the pKa of the conjugate acid is **stronger** than the acid, so the equilibrium is shifted to the right. Therefore, the right reaction is a stronger acid.

Choosing a Base to Deprotonate a Given Compound

Suppose you need to deprotonate the following compound: **acid**

Deprotonate means remove the proton (acid).

First, find the pKa value: **acid** = 10

Find the chemical equation of the conjugate base with a conjugate base (B).



The pKa of the conjugate acid must be higher than 10, so the reaction is favored.

Find a conjugate base with a pKa value of 10. Acids are weak acids (pKa > 10).



Choosing an Acid to Protonate a Given Compound

Suppose you need to protonate the following compound: **acid** (the acid is the acid).

Acid is a very reactive base and needs to be protonated before being deprotonated.

To do this, write the chemical equation of the acid with a weak acid.



The pKa of the conjugate acid of the acid is 10. Therefore, the pKa of the acid.

and the pKa of the acid is 10. Acids are weak acids (pKa > 10).

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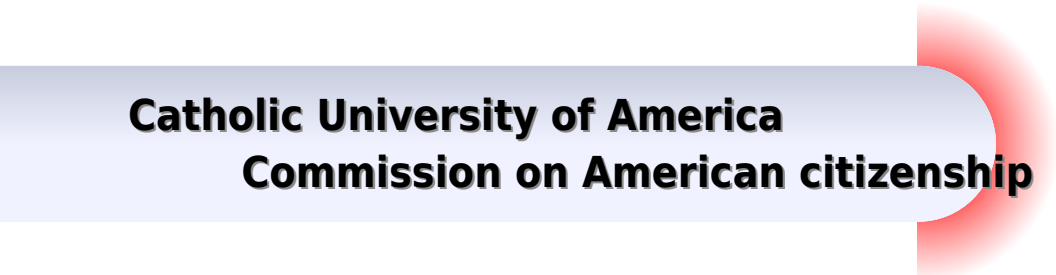
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