

## CHEM 102 – Recitation Week 7 – Answer Key

Ch. 16, 33. In each reaction, identify the Brønsted-Lowry acid, the Brønsted-Lowry base, the conjugate acid, and the conjugate base. Use symbols A, B, CA, and CB, respectively.

- $\text{H}_2\text{CO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HCO}_3^-(\text{aq})$
- $\text{NH}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq})$
- $\text{HNO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$
- $\text{C}_2\text{H}_5\text{N}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{C}_2\text{H}_5\text{NH}^+(\text{aq}) + \text{OH}^-(\text{aq})$

### Concepts needed:

Brønsted-Lowry definition of acids and bases:

definition of conjugate acid-base pairs: their chemical formulas must differ by one proton,  $\text{H}^+$ .

**Note:** proton = an atom of hydrogen that lost its electron, hence the  $\text{H}^+$  symbol.  
chemical equations

### Answer:



**Note:** Brønsted-Lowry bases must have a lone pair of electrons available to bind the incoming proton by a covalent bond in which both shared electrons come from the base.

Recall from CHEM 101 how to determine the number of lone pairs in atoms of non-metals!  
Such lone pairs are usually not shown in chemical formulas.



# Chem 102 Lab Manual Answer Key

**Rose Arny**

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